

47
YEARS
OF EXCELLENCE



NARAYANA
IIT-JEE/NEET/FOUNDATION

RE-NEET (UG) 2026



**PAPER &
SOLUTIONS**

PAPER CODE -70



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PHYSICS

1. A photon and an electron, each of 20 eV energy, move in free space. The ratio of linear momentum of electron p_e to that of photon p_{ph} , $\frac{p_e}{p_{ph}}$ is :

(Take speed of light = $3 \times 10^8 \text{ ms}^{-1}$, charge of electron = $-1.6 \times 10^{-19} \text{ C}$ and mass of electron = $9 \times 10^{-31} \text{ kg}$)

- (1) 275 (2) $\frac{2}{450}$
 (3) $\frac{1}{250}$ (4) 225

Ans. (4)

Sol. $P_e = \sqrt{2mE}$

$$P_p = \frac{E}{C}$$

$$\frac{P_e}{P_h} = \frac{\sqrt{2mE}}{E} \times C$$

$$= \frac{\sqrt{2 \times 9 \times 10^{-31} \times 20 \times 1.6 \times 10^{-19}}}{20 \times 1.6 \times 10^{-19}} \times 3 \times 10^8$$

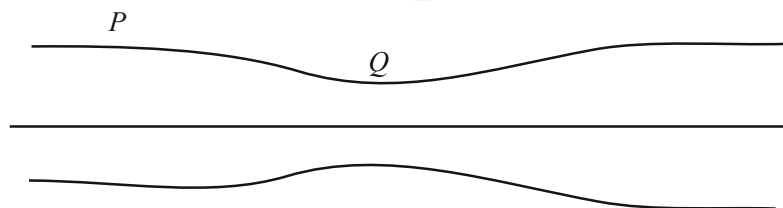
$$= \frac{3 \times 2 \times 4 \times 10^{-25}}{20 \times 16 \times 10^{-19}} \times 3 \times 10^8$$

$$= \frac{3}{4} \times 3 \times 10^{-6} \times 10^8$$

$$= \frac{900}{4} = 225$$

2. Water flows in a streamline motion through a horizontal pipe of circular cross-section as shown in the figure. The pressure difference of water between P and Q is 15 Nm^{-2} . The area of cross-section at P and Q are 40 cm^2 and 20 cm^2 , respectively. The rate of flow of water through the pipe, in $\text{cm}^3 \text{ s}^{-1}$, is :

[Take density of water = 1000 kg m^{-3}]



- (1) 400 (2) 100
 (3) 200 (4) 300

Ans. (1)

Sol. $\Delta P = \rho gh$

$$\frac{dQ}{dt} = \sqrt{\frac{2gh}{A_1^2 - A_2^2}} A_1 A_2$$

$$= \sqrt{\frac{2\Delta P}{\rho(A_1^2 - A_2^2)}} \times A_1 A_2$$

$$L = (\mu_0 n \cdot n l \pi r^2)$$

$$= (\mu_0 n^2 l) \pi r^2$$

5. The temperature of a metallic sphere of radius R is increased by a small amount ΔT . If the linear coefficient of thermal expansion of the metal is α , the approximate increase in the volume of the sphere is :

- (1) $6\pi R^3 \alpha \Delta T$ (2) $2\pi R^3 \alpha \Delta T$
 (3) $3\pi R^3 \alpha \Delta T$ (4) $4\pi R^3 \alpha \Delta T$

Ans. (4)

Sol. $V = \frac{4}{3} \pi R^3$

$$R_R = R(1 + 2\Delta T)$$

$$V_R = \frac{4}{3} \pi R^3 (1 + \alpha \Delta T)^3$$

$$V_R = \frac{4}{3} \pi R^3 (1 + 3\alpha \Delta T)$$

$$\Delta V = \frac{4}{3} \pi 3\alpha \Delta T R^3$$

$$\Delta V = 4\pi \alpha \Delta T R^3$$

'OR'

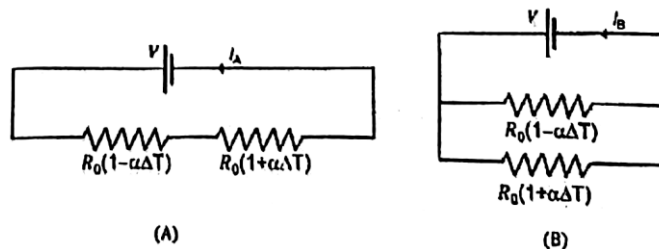
$$V = \frac{4}{3} \pi R^3$$

$$\Delta V = \frac{4}{3} \pi 3R^2 \Delta R$$

$$\Delta V = \left(\frac{4}{3} \pi 3R^2 R \alpha \Delta T \right)$$

$$= (4\pi R^3) \alpha \Delta T$$

6. Consider two circuits, (A) and (B), each having two resistors. One of them has a positive temperature coefficient of resistance, $+\alpha$, while the other one has a negative temperature of coefficient, $-\alpha$, as shown in the figure. The current through these circuits are denoted by I_A and I_B . At initial temperature, the resistance of the two resistors is R_0 . As the temperature is increased, the correct option that describes the variation of current in these circuits is :



- (1) both I_A and I_B remain constant (2) I_A remains constant while I_B increases
 (3) I_A decreases while I_B increases (4) I_A increases while I_B decreases

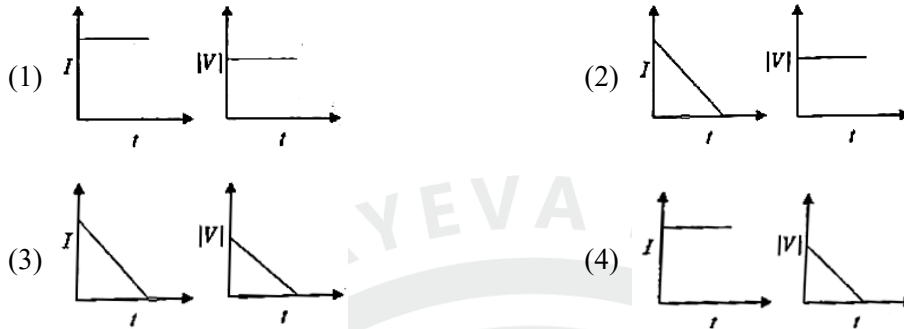
Ans. (2)

Sol. $R_S = +2R_0$

$$R_p = \frac{R_0(1 - \alpha\Delta T)R_0(1 + \alpha\Delta T)}{2R_0}$$

$$R_p = \frac{R_0(1 - \alpha^2\Delta T^2)}{2}$$

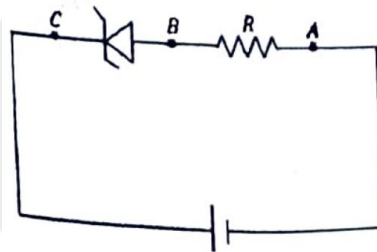
7. A beam of light falls on a metal surface such that photo-electrons are generated. If power of the light source starts to decrease linearly with time t , then variation of the photocurrent I and magnitude of the stopping potential $|V|$ with time is best represented by :



Ans. (2)

Sol. $P = \frac{N h \nu}{t}$

8. An ideal Zener diode with breakdown voltage of -3 V is reverse biased with a negative input voltage $V_i = -5\text{ V}$. The magnitude of voltage difference between points B and A is :

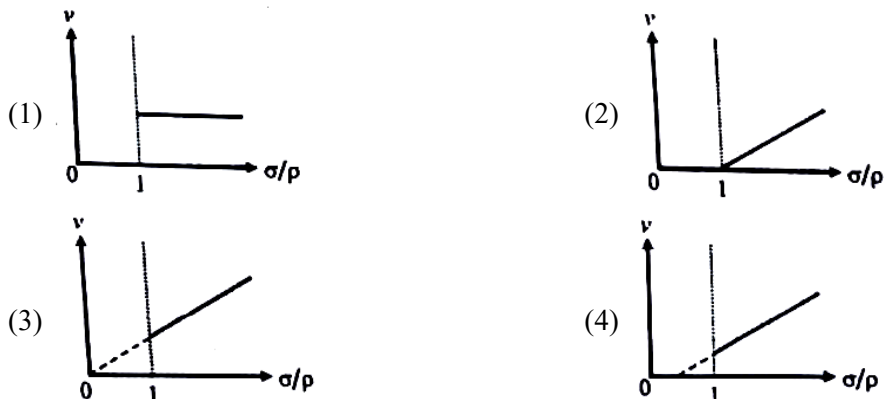


- (1) 0 V (2) 3 V
 (3) 2 V (4) 1 V

Ans. (3)

Sol. $V = 5 - 3 = 2\text{ V}$

9. In the measurement of viscosity of liquids using terminal velocity experiment, spherical balls of same radius but having different densities are used. The variation of the terminal velocity (v) with the ratio of density of spherical ball (σ) to density of the liquid (ρ), is best represented by:



Ans. (2)

Sol.
$$v = \frac{2(\sigma - \rho)r^2}{9\eta} \cdot g$$

$$v = \frac{2}{9}\rho \left[\frac{\sigma - 1}{\rho} \right] g$$

$$y = m(x - 1)$$

$$y = mx - m.$$

10. Two planets P_1 and P_2 with equal mass have radii R_1 and R_2 , respectively, where $R_2 = \frac{R_1}{2}$. The

escape speeds of P_1 and P_2 are v_1 and v_2 , respectively. Then $\frac{v_2}{v_1}$ is :

(1) 2

(2) $\frac{1}{\sqrt{2}}$

(3) 1

(4) $\sqrt{2}$

Ans. (4)

Sol.
$$R_2 = \frac{R_1}{2}$$

$$v_1 = \sqrt{\frac{2GM}{R_1}}$$

$$v_2 = \sqrt{\frac{2GM}{R_2}}$$

$$\frac{v_2}{v_1} = \sqrt{\frac{R_1}{R_2}}$$

$$\frac{v_2}{v_1} = \sqrt{\frac{2}{1}}$$

$$\left[\frac{v_2}{v_1} = \sqrt{2} \right]$$

11. An ac voltage $V = 220 \sin(2 \times 10^3 t)$ Volt is applied to a series LCR circuit. Then the current amplitude in this circuit is : (Given : $L = 10$ mH, $C = 25$ μ F, $R = 100$ Ω)

(1) 22.0 A

(2) 2.2 A

(3) 5.5 A

(4) 11.0 A

Ans. (2)

Sol.
$$x_C = \frac{1}{W_C} = 25x$$

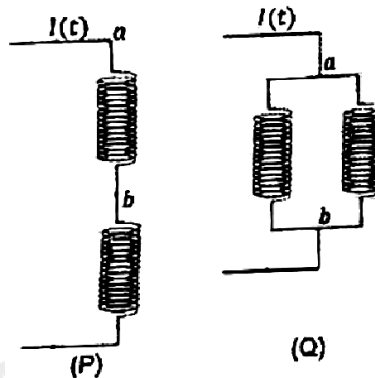
$$= \frac{1}{2 \times 10^3 \times 25 \times 10^{-6}} \Rightarrow \frac{10^3}{50} = \frac{1000}{50}$$

$$x_L = W_L = 2 \times 10^3 \times 10^{-3} \times 10^{20} = 20$$

$$x_L = x_C \rightarrow \text{resonance}$$

$$I_0 = \frac{V_0}{R} = \frac{220}{100} = 2.2 \text{ A}$$

12. Two identical inductors are connected in two different configurations P and Q, where a time varying current $I(t)$ is flowing, as shown in the figure. The induced emf between points a and b for configuration P is E_p and that for configuration Q is E_Q . The ratio E_p / E_Q is :
[Neglect the effect of mutual inductance.]

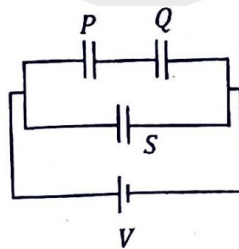


- (1) 2
(2) 1/4
(3) 1/2
(4) 1

Ans. (1)

Sol. $E_p = L \frac{dI}{dt}$
 $E_Q = \frac{L dI}{2 dt} = \frac{L dI}{2 dt}$
 $\frac{E_p}{E_Q} = \frac{2}{1}$

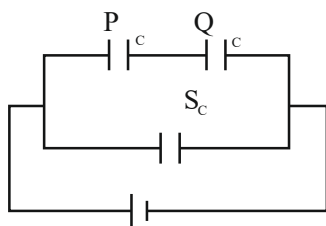
13. Three identical capacitors P, Q and S, each of the capacitance C , are connected to a battery of voltage V , as shown in the figure. If the energy stored in the capacitor P and total energy stored in the system are U_P and U_T , respectively, then the ratio $\frac{U_P}{U_T}$ is :



- (1) 1/6
(2) 2/3
(3) 1/3
(4) 1/2

Ans. (1)

Sol.



$$U_P = \frac{1}{2} C \left(\frac{V}{2} \right)^2$$

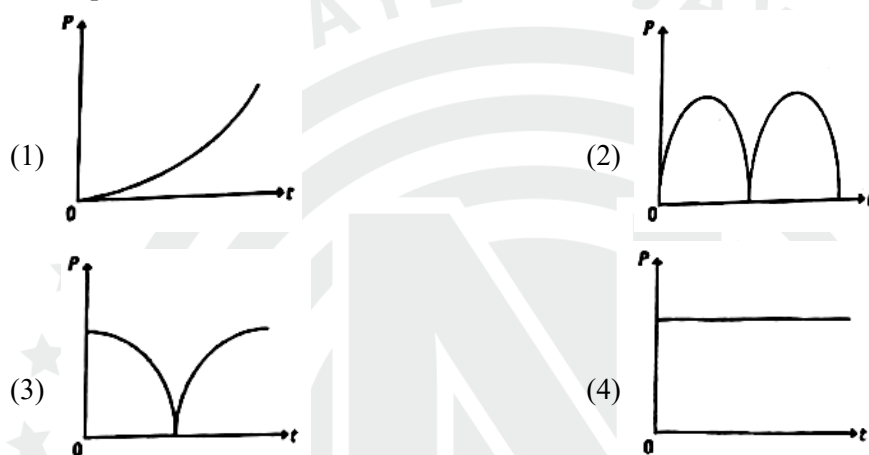
$$= \frac{1}{2} C \frac{V^2}{4}$$

$$U_T = \frac{1}{2} \times \frac{3}{2} CV^2$$

$$\frac{U_P}{U_T} = \frac{1}{4} \times \frac{2}{4} \times \frac{1}{3}$$

$$= \frac{1}{6}$$

14. A conducting loop of finite resistance lies on the x - y plane. There is a constant magnetic field in the z direction. The area of the loop varies with time t , as $A = A_0(1 + \sin t)$ in appropriate units. The figure that correctly indicates the qualitative behaviour of the power P dissipated in the loop as a function of time is :



Ans. (3)

Sol. $\phi = BA$

$$e = \frac{d\phi}{dt} = B \cdot \frac{dA}{dt}$$

$$= B_0 A_0 \cos t$$

$$P = \frac{e^2}{R}$$

$$= B^2 A_0^2 \cos^2 t$$

$$P = B^2 A_0^2 (1 + \cos 2\omega t)$$

15. In an adiabatic expansion, the temperature of one mole of an ideal monatomic gas ($\gamma = 5/3$) decreases from 60 K to 50 K. The work done by the gas in the process is :

(Take the universal gas constant as $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$)

(1) 166 J

(2) 41.5 J

(3) 83 J

(4) 124.5 J

Ans. (4)

Sol. $W = Q$

$$= n c_v \Delta T = 1 \times \frac{3}{2} R (10)$$

$$= \frac{3}{5} \times \frac{25}{5} \times 10 = 125$$

16. Consider a particle moving along a straight line, whose position as a function of time is given by $s(t) = \alpha t^2 - \beta t + \gamma$, where $\alpha = 1 \text{ ms}^{-2}$, $\beta = 6 \text{ ms}^{-1}$ and $\gamma = 5 \text{ m}$. The average speed of the particle, in ms^{-1} , from $t = 0$ to $t = 6 \text{ s}$ is :

- (1) 0 (2) 12
(3) 6 (4) 3

Ans. (4)

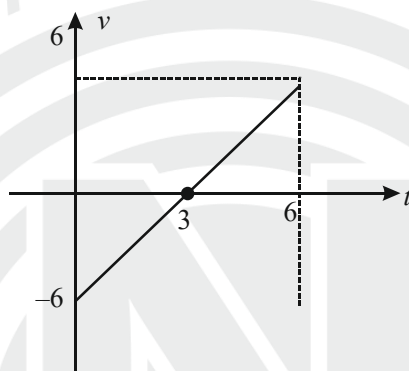
Sol. $s = \alpha t^2 - \beta t + \gamma$

$$s = t^2 - 6t + 5$$

$$v = 2t - 6$$

$$s = \frac{1}{2} \times 3 \times 6 \times 2 = 18$$

$$v = \frac{18}{6} = 3$$



17. The following table presents the part of the electromagnetic spectrum and their corresponding major applications.

| Part of the electromagnetic spectrum | | Applications | |
|--------------------------------------|------------|--------------|-------------------------------------|
| P | Microwave | I | For purifying the water |
| Q | UV rays | II | For warming the food |
| R | Gamma rays | III | For AM and FM communication systems |
| S | Radio wave | IV | For treating the Cancer cells |

The correct option is :

- (1) P-II, Q-IV, R-III, S-I (2) P-I, Q-II, R-III, S-IV
(3) P-I, Q-IV, R-II, S-III (4) P-II, Q-I, R-IV, S-III

Ans. (4)

Sol.

| Part of the electromagnetic spectrum | | Applications | |
|--------------------------------------|------------|--------------|-------------------------------------|
| P | Microwave | II | For warming the food |
| Q | UV rays | I | For purifying the water |
| R | Gamma rays | IV | For treating the Cancer cells |
| S | Radio wave | III | For AM and FM communication systems |

18. An ideal gas is made of polyatomic molecules. Each of the molecules has three translational, three rotational and f number of vibrational modes. If the ratio of heat capacities C_p / C_v of the gas is $8 / 7$, then the value of f is :

- (1) 1 (2) 4
(3) 3 (4) 2

Ans. (2)

Sol. 3 translation + 3 rotational + f vibrational

$$\gamma = \frac{8}{7} = 1 + \frac{2}{7}$$

$$\frac{8}{7} = 1 + \frac{2}{f}$$

$$\frac{1}{7} = \frac{2}{f}$$

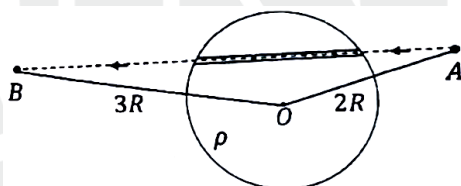
$$f = 14$$

$$3t + 3r + 8f$$

4 kinetic and 4 potential vibrational mode.

19. A unit positive point charge is taken slowly through an infinitesimally thin tube that is inside a charged dielectric sphere of radius R , having uniform positive charge density ρ , as shown in the figure. The initial and final positions of the charge are marked by A and B at distances $2R$ and $3R$ respectively, from the centre of the sphere. In this process, the magnitude of the total work done on the point charge is $\frac{\rho R^2}{n\epsilon_0}$. The value of n is :

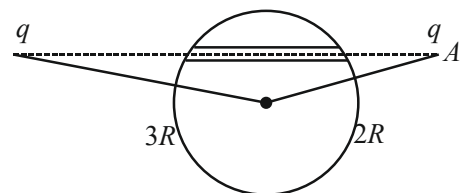
(ϵ_0 is the permittivity of vacuum)



- (1) 18 (2) 2
(3) 6 (4) 9

Ans. (1)

Sol.



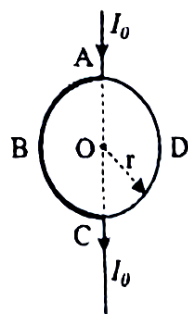
$$U_A = \frac{KQq}{2R}, \quad U_B = \frac{KQq}{3R}$$

$$\Delta U = \frac{KQq}{R} \left(\frac{1}{6} \right)$$

$$= \frac{1}{4\pi\epsilon_0} \frac{q\rho}{R} \frac{4}{3} \pi R^3 \frac{1}{6}$$

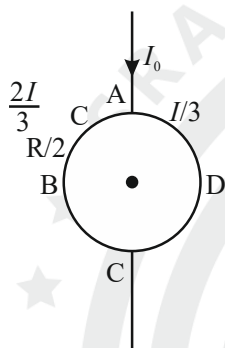
$$\Rightarrow \frac{q\rho R^2}{18\epsilon_0}$$

20. A current I_0 flows through a metallic circular loop of radius r as shown in the figure. Resistance of the segment ABC is half that of ADC. Magnitude of magnetic field at the center O of the loop is :



- (1) $\frac{\mu_0 I_0}{2\pi r}$ (2) $\frac{\mu_0 I_0}{12r}$ (3) $\frac{\mu_0 I_0}{4r}$ (4) $\frac{\mu_0 I_0}{2r}$

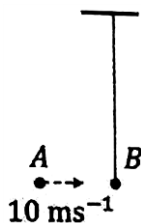
Ans. (2)
Sol.



B_{net} at centre

$$\begin{aligned} &= \frac{\mu_0 2I}{2a3} \times \frac{1}{2} - \frac{\mu_0}{2a} \times \frac{I}{3} \times \frac{1}{2} \\ &= \frac{\mu_0}{2a} \times \frac{I}{3} \times \frac{1}{2} [2-1] \\ &= \frac{\mu_0 I}{1\pi a} \end{aligned}$$

21. Bob B of mass m at rest is hanging vertically from the ceiling via a massless string of length 10 m, as shown in the figure. Point mass A of mass m travelling horizontally with speed 10 ms^{-1} hits bob B elastically. The bob B rises h meter after the collision. Taking the acceleration due to gravity $g = 10 \text{ ms}^{-2}$ and neglecting the size of the bob, the value of h is :



- (1) 2.5 (2) 8 (3) 7 (4) 5

Ans. (4)

Sol. $\frac{1}{2} m(10)^2 = mgh$

$$h = \frac{100}{20} = 5 \text{ m}$$

22. An electromagnetic wave travelling in a lossless dielectric medium having a dielectric constant, $\epsilon_r = 9$, has the electric field, $E_x = E_0 \sin(kz - 2\pi \times 10^6 t)$ Vm^{-1} where E_0 is the amplitude and k is the wave vector. Among the following options, the incorrect choice is :
- (1) The direction of propagation of the electromagnetic wave is along $+z$
 - (2) The speed of the electromagnetic wave inside the medium is 10^8 ms^{-1}
 - (3) The wavelength of the electromagnetic wave inside the medium is 300 m
 - (4) The magnetic field is given by the relation $B_y = \frac{B_0}{v} \sin(kz - 2\pi \times 10^6 t)$ where v is the speed of the electromagnetic wave inside the medium

Ans. (3)

Sol. direction of propagation -Z-axis

$$V = \frac{c}{\sqrt{\epsilon_r}} \Rightarrow \frac{3 \times 10^8}{3} = 10^8 \text{ m/s}$$

$$\text{wavelength } \lambda = \frac{v}{f} = \frac{10^8}{10^6} = 100 \text{ m}$$

23. A particle of mass M moves along a horizontal x axis from $x = 0$ to $x = L$. The coefficient of kinetic friction varies as a function of x as $\mu_k(x) = \mu_0 - \alpha x$, where μ_0, α are constants of appropriate dimensions, so that $\mu_k(L) = 0$. The total work done by the frictional force during the motion is $n\mu_0 MgL$, where g is the acceleration due to gravity. The value of n is :

- (1) $\frac{1}{2}$
- (2) 3
- (3) 1
- (4) $\frac{1}{3}$

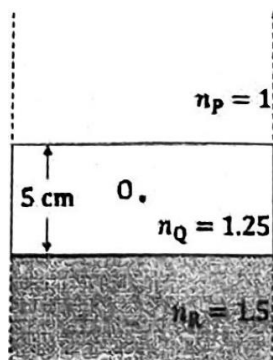
Ans. (1)

Sol. $W = mg \int_0^L (\mu_0 - \alpha x) dx$

$$= mg \left[\mu_0 L - \frac{\alpha L^2}{2} \right] = mg \left[\mu_0 L - \frac{\mu_0 L}{2} \right]$$

$$n \times \mu_0 mgL = mg \left[\frac{\mu_0 L}{2} \right] \quad n = \frac{1}{2}$$

24. Consider three media P, Q and R with refractive indices 1, 1.25, and 1.5, respectively. The medium Q having a thickness of 5 cm is placed between extended media P and R as shown in the figure. An object O is placed at the center of medium Q. If viewed from medium P near the normal direction, the apparent depth of O is h_1 . For similar observation from medium R, the apparent depth is h_2 . The value of $|h_1 - h_2|$, in cm, is :



(1) 3

(2) 0

(3) 1

(4) 2

Ans. (3)

Sol. $n_p = 1$ $h_1 = h \times \frac{n_p}{n_Q}$

$n_Q = 1.25$ $h_1 = 2.5 \times \frac{1}{1.25} = 2 \text{ cm}$

$n_R = 1.5$ $h_2 = d \frac{n_R}{n_Q} = 2.5 \times \frac{1.5}{1.25} = 3 \text{ cm}$

$|h_1 - h_2| = 1 \text{ cm}$

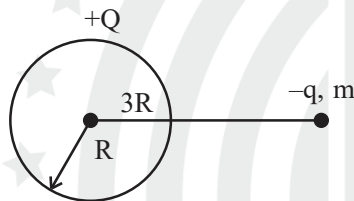
25. Consider a fixed uniformly charged insulating sphere with radius R and total charge $+Q$. A point charge $-q$ ($q \ll Q$) with mass m is released from rest at a distance of $3R$ from the centre of the charged sphere. When the point charge reaches the surface of the sphere, its speed is :

(ϵ_0 is the permittivity of vacuum, neglect gravitational forces.)

- (1) $\sqrt{\frac{Qq}{4\pi\epsilon_0 mR}}$ (2) $\sqrt{\frac{3Qq}{4\pi\epsilon_0 mR}}$
- (3) $\sqrt{\frac{2Qq}{3\pi\epsilon_0 mR}}$ (4) $\sqrt{\frac{Qq}{3\pi\epsilon_0 mR}}$

Ans. (4)

Sol.



$$0 + (-q) \left[\frac{KQ}{3R} \right] = \frac{1}{2}mv^2 - q \times \frac{KQ}{R}$$

$$\frac{1}{2}mv^2 = \frac{KQq}{R} - \frac{KQq}{3R} \Rightarrow \frac{2KQq}{3R}$$

$$v \Rightarrow \sqrt{\frac{4KQq}{3mR}} \Rightarrow \sqrt{\frac{Qq}{3\pi\epsilon_0 mR}}$$

26. A car travels on a circular racetrack of radius 50 m , which is banked at an angle θ . If the car travels at a speed 10 ms^{-1} , then the wear and tear on its tyres is minimum. Taking the acceleration due to gravity to be 10 ms^{-2} , the value of θ is :

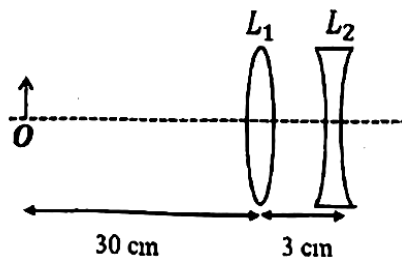
- (1) $\tan^{-1}(2\sqrt{3})$ (2) $\tan^{-1}\left(\frac{1}{5}\right)$
- (3) $\tan^{-1}\left(\frac{2}{5}\right)$ (4) $\tan^{-1}(\sqrt{3}/2)$

Ans. (2)

Sol. $\tan \theta = \frac{v^2}{Rg}$

$\theta = \tan^{-1}\left(\frac{1}{5}\right)$

29. The lens combination as shown in the figure, consists of two lenses, L_1 and L_2 , of the focal lengths +10 cm and -10 cm, respectively. The position of the image formed is :



- (1) 60 cm to the right of the concave lens
 (2) 20 cm to the left of the concave lens
 (3) 60 cm to the left of the concave lens
 (4) 30 cm to the right of the concave lens

Ans. (3)

Sol. $f_1 = 10$ cm

$f_2 = -10$ cm

(1) for convex lens

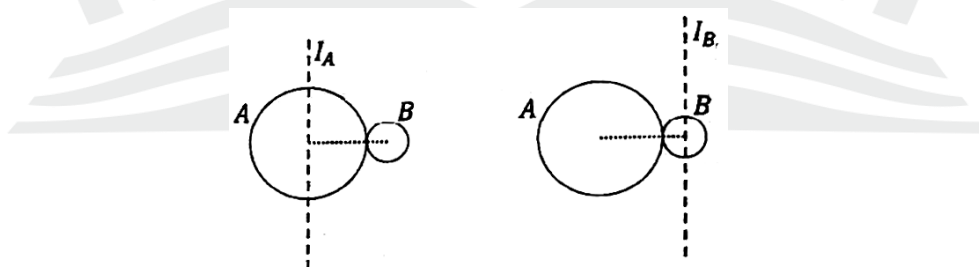
$$\frac{1}{10} = \frac{1}{v} - \frac{1}{(-30)}, v = 15 \text{ cm}$$

(2) for concave lens object distance is 12 cm hence

$$\frac{1}{-10} = \frac{1}{v} - \frac{1}{(12)}$$

$v = 60$ cm

30. A solid sphere A of radius R and mass M is attached at a point to a smaller solid sphere B of radius $r < R$ and mass $m < M$. Assume that the line joining their centres lies along the horizontal. The moment of inertia of the system calculated about a vertical axis passing through the centre of A is I_A and that calculated about a vertical axis passing through the centre of B is I_B . The difference $I_A - I_B$ is :



(1) 0

(2) $(M - m)(R + r)^2$

(3) $(m - M)(R + r)^2$

(4) $(m - M)(R - r)^2$

Ans. (3)

Sol. $I_A = \frac{2}{5}MR^2 + \left[\frac{2}{5}mR^2 + m(R + r)^2 \right]$

$$I_B = \frac{2}{5}mR^2 + \left[\frac{2}{5}MR^2 + M(R + r)^2 \right]$$

$$I_A - I_B = (m - M)(R + r)^2$$

31. Consider that an electron is revolving in an excited state of Hydrogen atom with velocity $\sqrt{25.6} \times 10^5 \text{ ms}^{-1}$. The radius of the orbit is $x \times 10^{-9} \text{ m}$. The value of x is :
[Take the mass of electron to be $9 \times 10^{-31} \text{ kg}$, charge of electron

$$= -1.6 \times 10^{-19} \text{ C and } \frac{1}{4\pi\epsilon_0} = 9 \times 10^9 \text{ N m}^2 \text{ C}^{-2}]$$

- (1) 1 (2) 4 (3) 3 (4) 2

Ans. (1)

Sol. $V = \sqrt{25.6} \times 10^5$

$$m = 9 \times 10^{-31}$$

$$e = 1.6 \times 10^{-19}$$

$$\frac{mv^2}{r} = \frac{1}{4\pi\epsilon_0} \frac{e^2}{r^2}$$

$$r = \frac{1}{4\pi\epsilon_0} \frac{e^2}{mv^2}$$

$$r = \frac{9 \times 10^9 \times (1.6 \times 10^{-19})^2}{9 \times 10^{-31} \times (25.6 \times 10^{10})}$$

$$r = x \times 10^{-9}$$

$$x = 1$$

32. The mean free path of molecules in an ideal gas A is half that of another ideal gas B. The diameter of the spherical molecules of gas A is twice the diameter of the molecules of B. If number densities of the gases A and B are n_A and n_B , respectively, then the correct option is :

- (1) $n_A = \frac{1}{2} n_B$ (2) $n_A = n_B$ (3) $n_A = 2n_B$ (4) $n_A = \frac{1}{4} n_B$

Ans. (1)

Sol. $\lambda_A = \frac{1}{2} \lambda_B$ $d_A = 2d_B$

$$\lambda = \frac{1}{\sqrt{2}\pi d^2 n} \quad \frac{\lambda_A}{\lambda_B} = \frac{d_B^2 n_B}{d_A^2 n_A} \Rightarrow \frac{1}{2} = \frac{1}{4} \times \frac{n_B}{n_A} = \frac{n_B}{n_A} = 2$$

33. A cylindrical cork of uniform density floats in a liquid of density ρ_1 . If the cork is depressed slightly and released, it oscillates harmonically with time period T . If the same cork floats in another liquid of density ρ_2 , then the similar oscillation has time period $2T$. The value of ρ_2 / ρ_1 is :

- (1) 1/4 (2) 4
(3) 2 (4) 1/2

Ans. (1)

Sol. $T \propto \frac{1}{\sqrt{\rho}}$

$$T_2 = 2T_1$$

$$\frac{T_2}{T_1} = 2 = \sqrt{\frac{\rho_1}{\rho_2}}$$

$$\frac{\rho_2}{\rho_1} = \frac{1}{4}$$

34. For sound waves, if the number of nodes for the 5th harmonic of an open-ended pipe is n and that for the 9th harmonic of the same pipe with one of its ends closed is m , the ratio $\frac{n}{m}$ is :

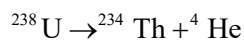
- (1) $\frac{3}{5}$ (2) $\frac{5}{9}$ (3) $\frac{9}{5}$ (4) 1

Ans. (4)

Sol. $n = 5, m = 5$ (For open organ pipe number of nodes $n = 5$) $\frac{n}{m} = 1$

(For close organ pipe number of nodes $m = 5$)

35. Consider the following nuclear reaction :



Take masses of ${}^{238}\text{U}, {}^{234}\text{Th}$ and ${}^4\text{He}$ as 238.050u, 234.043u and 4.003 u , respectively. The Q value for the reaction, in keV , is :

[Given : $1\text{u} = 931.5 \text{ MeVc}^{-2}$]

- (1) 3740 (2) 3726 (3) 3730 (4) 3736

Ans. (2)

Sol. ${}^{238}\text{U} \rightarrow {}^{234}\text{Th} + {}^4\text{He}$

$$\Delta m \Rightarrow 238.050 - (234.043 + 4.003)$$

$$\Delta m = 0.004 \times 931.5$$

$$= 3.726$$

$$= 3726$$

36. Which of the following measurements require 'index correction' ?

- (1) Measurement of speed of sound Using resonance tube
 (2) Measurement of resistance of a wire using meter bridge
 (3) Measurement of gravitational acceleration using simple Pendulum
 (4) Measurement of focal length of lenses using optical bench

Ans. (4)

Sol. Practical Physics

37. In a solar system, the time-period revolution of a planet tracing a circuit orbit of radius R is proportional to :

- (1) R^3 (2) $R^{1/2}$ (3) $R^{3/2}$ (4) R^2

Ans. (3)

Sol. $T \propto R^{3/2}$

38. Consider that σ_s, k_B, b represent Stefan-Boltzmann constant, Boltzmann constant and Wien's displacement law constant, respectively. The dimension of $\sigma_s k_B^{-1} b$ is :

- (1) $[L^{-1} T^{-1} K^{-4}]$ (2) $[L^{-1} T^{-1} K^{-2}]$ (3) $[L^{-1} K^{-2}]$ (4) $[L^{-1} T^{-1} K^{-3}]$

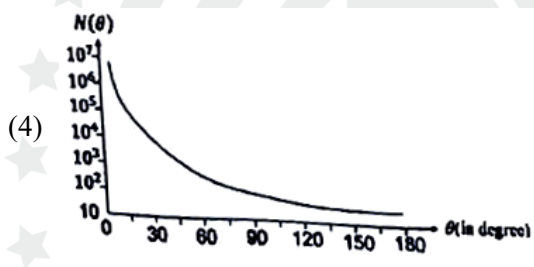
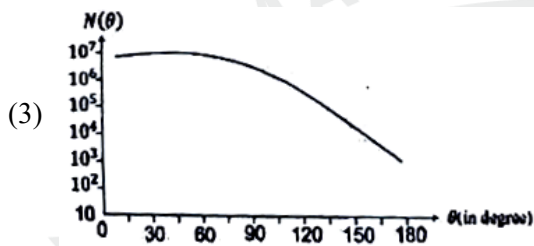
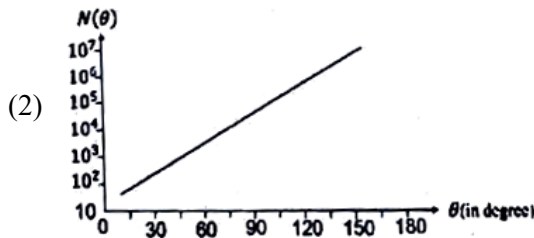
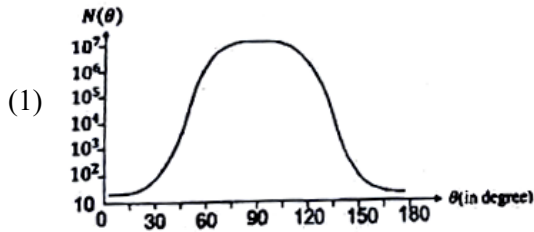
Ans. (2)

Sol. $[\sigma_B] = [NT^{-3}R^{-4}]$

$$[k_B] = [11L^2F^{-2}K^{-1}]$$

$$\sigma_{\Delta} K_B^{-1} b = [f / T^{-3} K^{-4}] [y^{-1} L^2 I^2 K'] [LK] = [L^1 T^{-1} K^{-2}]$$

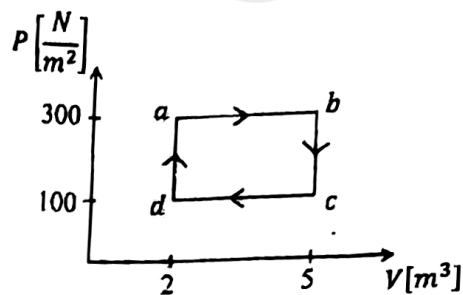
42. In Geiger-Marsden experiment, the number of scattered α -particles $N(\theta)$ is plotted as a function of scattering angle θ . Which of the following options represents the correct plot ?



Ans. (4)

Sol. $N \propto \frac{1}{\sin^4(\theta/2)}$

43. One mole of an ideal monatomic gas undergoes a cyclic process as shown in the figure. The total heat supplied to the gas is :



- (1) 800 J (2) 400 J
(3) 500 J (4) 600 J

Ans. (4)

Sol. $\Delta Q = \Delta W$

$\Delta Q = 200 \times 3 = 600 \text{ J}$

44. Two infinitely long parallel conducting wires A and B carry currents I and $2I$, respectively, in the same direction. The wire A has uniform mass per unit length λ and lies on an insulated floor. The wire B is kept fixed at a height h above the floor. The minimum magnitude of h so that the wire A does not rise from the floor is :

[g is the acceleration due to gravity and μ_0 is the permeability of free space.]

(1) $\frac{4\mu_0 I^2}{\pi\lambda g}$

(2) $\frac{\mu_0 I^2}{2\pi\lambda g}$

(3) $\frac{\mu_0 I^2}{\pi\lambda g}$

(4) $\frac{2\mu_0 I^2}{\pi\lambda g}$

Ans. (3)

Sol. $F = \frac{2\mu_0 II}{2\pi h} l = \frac{\mu_0 I^2 l}{\pi h}$

$$\frac{\mu_0 I^2 l}{\pi h} = \lambda l g$$

$$h = \frac{\mu_0 I^2}{\pi\lambda g}$$

45. Consider a spring-mass simple harmonic oscillator in one dimension. The mass of the particle is m kg and the spring constant is k Nm^{-1} . At a given instant, the extension of the spring is x meter and the speed of the particle is v ms^{-1} . On the $x - v$ plane, if the graph of v as a function of x is a circle, then the correct option is :

(1) $k = \sqrt{m}$

(2) $k = \frac{1}{m}$

(3) $k = m$

(4) $k = m^2$

Ans. (3)

Sol. $v = \omega\sqrt{A^2 - x^2}$

$$V^2 = \omega^2 A^2 - \omega^2 x^2$$

$$V^2 + \omega^2 x^2 = \omega^2 A^2$$

$$\frac{v^2}{\omega^2 A^2} + \frac{x^2}{A^2} = 1$$

$$\omega^2 A^2 = A^2$$

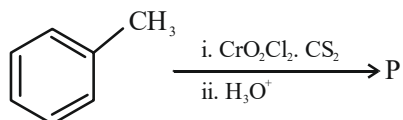
$$\omega^2 = 1$$

$$\frac{m}{k} = 1$$

$$m = k$$

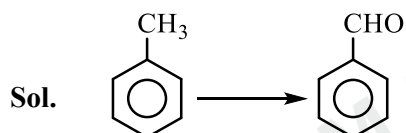
CHEMISTRY

46. Consider the following reaction, and chose the correct option.



- (1) Compound **P** is obtained by the hydrogenation of benzoyl chloride with Pd on BaSO_4 .
 (2) On treating compound **P** with saturated NaHCO_3 solution, brisk effervescence is observed.
 (3) Compound **P** can be prepared by treating benzene with anhydrous AlCl_3 and CH_3COCl
 (4) On treatment with bromine water, compound **P** gives a white precipitate.

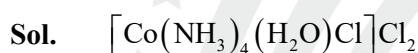
Ans. (1)



47. The formula of tetraammineaquachloridocobalt (III) chloride is

- (1) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}_2$ (2) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2] \times \text{H}_2\text{O}$
 (3) $[\text{Co}(\text{NH}_3)_4]\text{Cl}_3 \times \text{H}_2\text{O}$ (4) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}$

Ans. (1)



48. The lanthanide ion having four unpaired electrons is

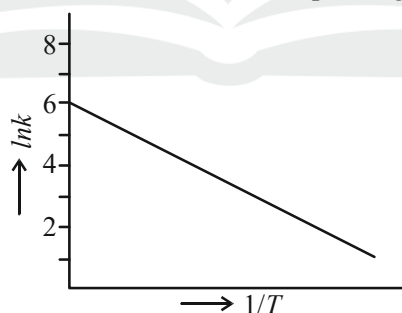
[Given: Atomic numbers of Ce = 58, Nd = 60, Tb = 65 and Ho = 67]

- (1) Ho^{3+} (2) Nd^{3+} (3) Ce^{3+} (4) Tb^{3+}

Ans. (1)



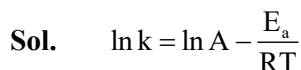
49. For an elementary chemical reaction, the Arrhenius plot is given below.



If the energy of activation is 6.64 kJ mol^{-1} and $R = 8.3 \text{ JK}^{-1} \text{ mol}^{-1}$, the temperature at which the rate constant becomes $\text{e}^2 \text{ min}^{-1}$, is

- (1) 250 K (2) 125 K (3) 150 K (4) 200 K

Ans. (4)



$$\ln k = 6 - \frac{E_a}{RT}$$

$$\ln e^2 = 6 - \frac{6.64 \times 10^3}{8.3 \times T}$$

$$2 \ln e = 4$$

$$= \frac{6.64 \times 10^3}{8.3 \times T}$$

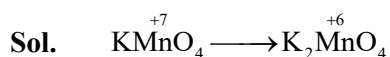
$$T = \frac{6.64 \times 10^3}{8.3 \times 4}$$

$$\boxed{T = 200\text{K}}$$

50. The green paramagnetic species formed by heating KMnO_4 at 513 K is

- (1) KO_2 (2) K_2MnO_4 (3) Mn_3O_4 (4) MnO

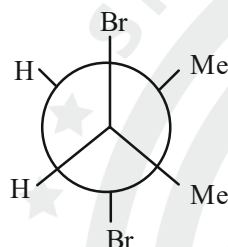
Ans. (2)



51. Given below are two statements:

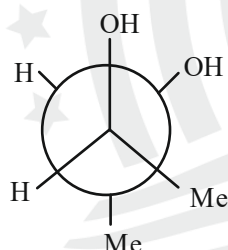
Statement I:

trans-But-2-ene upon treatment with Br_2 in CCl_4 gives the following product



Statement II:

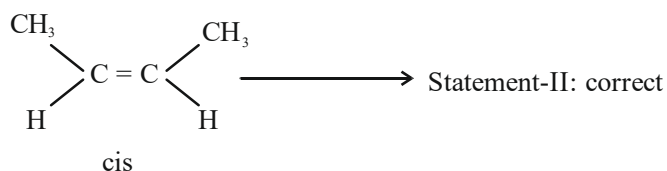
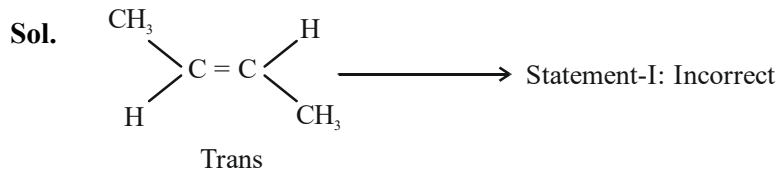
cis-But-2-ene upon treatment with alkaline KMnO_4 gives the following product



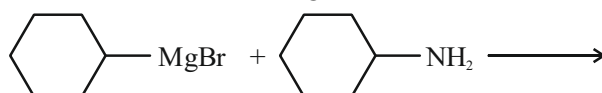
In the light of the above statements, choose the **most appropriate** answer from the options given below.

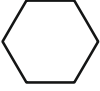
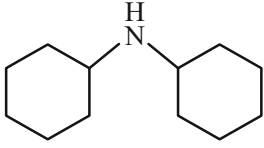
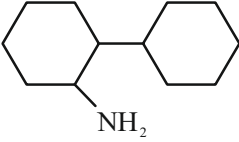
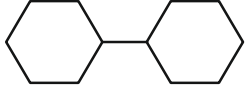
- (1) **Statement I** is incorrect but **Statement II** is correct
 (2) Both **Statement I** and **Statement II** are correct
 (3) Both **Statement I** and **Statement II** are incorrect
 (4) **Statement I** is correct but **Statement II** is incorrect

Ans. (1)

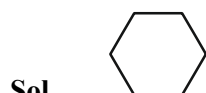


52. One of the products formed in the following reaction is



- (1) 
- (2) 
- (3) 
- (4) 

Ans. (1)



53. Given below are two statements:

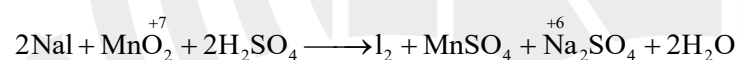
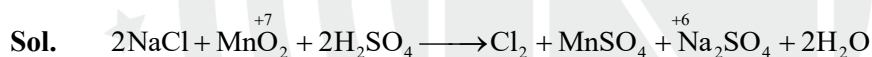
Statement-I: Heating NaCl with concentrated H_2SO_4 and MnO_2 results in oxidation of Mn.

Statement-II: Heating NaI with concentrated H_2SO_4 and MnO_2 results in reduction of Mn.

In light of the above statements, choose the **most appropriate** answer from the options given below:

- (1) Statement-I is incorrect but Statement-II is correct.
 (2) Both Statement-I and Statement-II are correct.
 (3) Both Statement-I and Statement-II are incorrect.
 (4) Statement-I is correct but Statement-II is incorrect.

Ans. (1)



54. Among the following options, the correct trend in the electron gain enthalpy is

- (1) $\text{I} > \text{Br} > \text{Cl} > \text{F}$ (2) $\text{F} > \text{Cl} > \text{Br} > \text{I}$
 (3) $\text{Br} > \text{Cl} > \text{F} > \text{I}$ (4) $\text{Cl} > \text{F} > \text{Br} > \text{I}$

Ans. (4)

Sol. $\text{Cl} > \text{F} > \text{Br} > \text{I}$

55. Given below are two statements:

Statement-I: $[\text{Fe}(\text{ox})_3]^{3-}$ is chiral.

Statement-II: $\text{trans-}[\text{Cr}(\text{H}_2\text{O})_2(\text{ox})_2]\Gamma$ is chiral.

In light of the above statements, choose the most appropriate answer from the options given below:

- (1) **Statement-I** is incorrect but **Statement-II** is correct.
 (2) Both **Statement-I** and **Statement-II** are correct.
 (3) Both **Statement-I** and **Statement II** are incorrect.
 (4) **Statement-I** is correct but **Statement-II** is incorrect.

Ans. (4)

56. The correct statement about peptides and proteins is
- (1) In α -helices, the polypeptide chain is twisted into a left-handed screw (helix) through intramolecular hydrogen bonds.
 - (2) Tertiary structure of proteins has two or more polypeptide subunits.
 - (3) Only the proteins having a quaternary structure are biologically active.
 - (4) In β -pleated sheet structures, peptide chains are held together by intermolecular hydrogen bonds.

Ans. (4)

Sol. Fact

57. Given below are two statements:

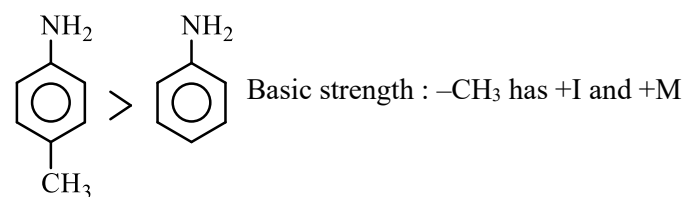
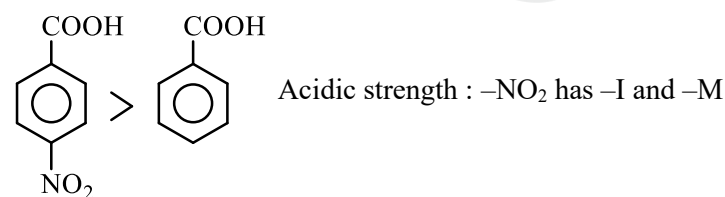
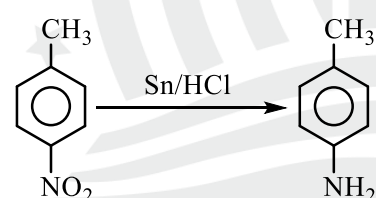
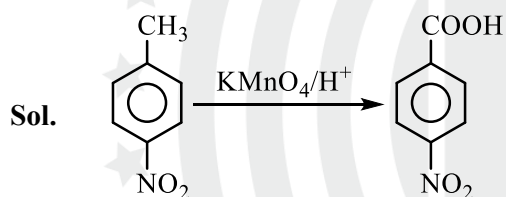
Statement-I: Oxidation of *p* nitrotoluene with acidic KMnO_4 gives an acid that is stronger than benzoic acid.

Statement-II: Reduction of *p* nitrotoluene with Sn/HCl followed by neutralization gives an amine that is more basic than aniline.

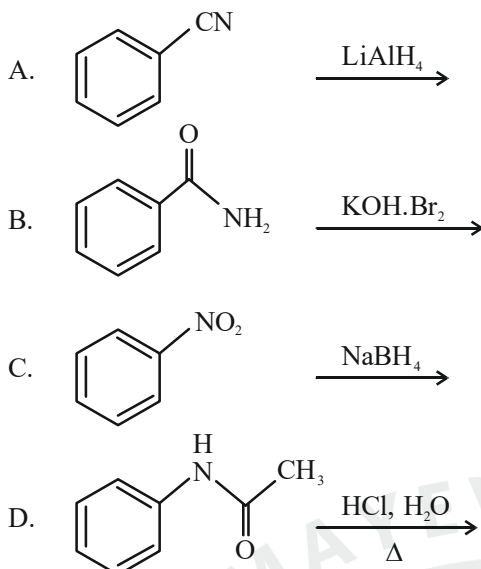
In light of the above statements, choose the most appropriate answer from the options given below.

- (1) Statement-I is incorrect but Statement-II is correct.
- (2) Both Statement-I and Statement-II are correct.
- (3) Both Statement-I and Statement-II are incorrect.
- (4) Statement-I is correct but Statement-II is incorrect.

Ans. (2)



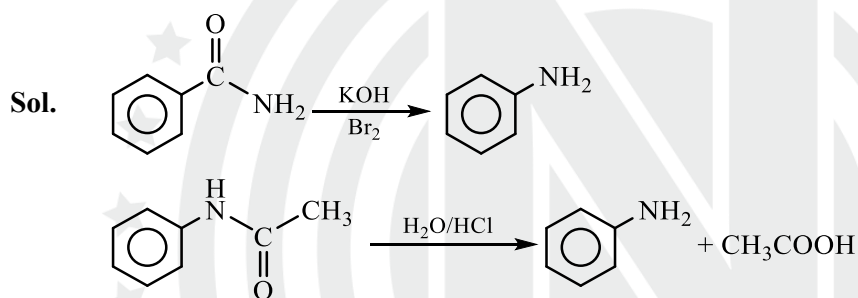
58. Identify the reactions which give aniline as the major product.



Choose the correct answer from the options given below.

- (1) C and D only (2) A and B only
 (3) B and D only (4) A and C only

Ans. (3)

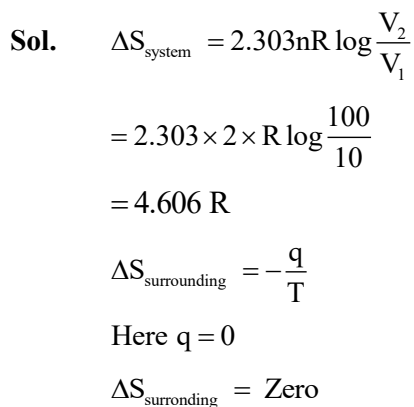


59. Two moles of an ideal gas undergo free expansion from 10 L to 100 L at 300 K. The values of ΔS_{system} and $\Delta S_{\text{surroundings}}$ are

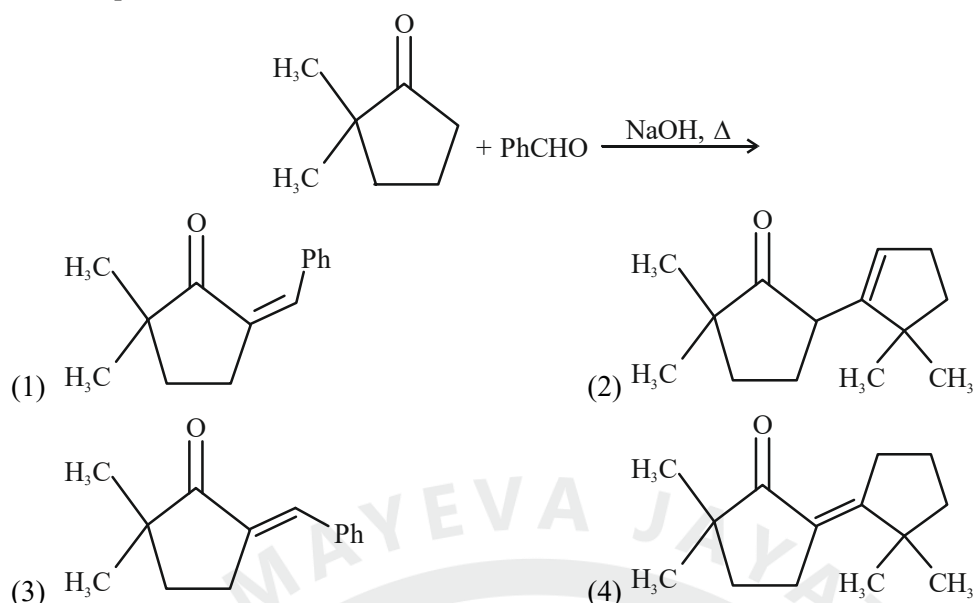
(R is universal gas constant)

- (1) $\Delta S_{\text{system}} = 4.606 R$; $\Delta S_{\text{surroundings}} = 0$
 (2) $\Delta S_{\text{system}} = 0$; $\Delta S_{\text{surroundings}} = 0$
 (3) $\Delta S_{\text{system}} = 4.606R$; $\Delta S_{\text{surroundings}} = -4.606R$
 (4) $\Delta S_{\text{system}} = 0$; $\Delta S_{\text{surroundings}} = 4.606R$

Ans. (1)



60. The compound that **CANNOT** be obtained from aldol condensation reaction shown below, is



Ans. (2)

Sol. Since position of double bond increasing is inside ring.

61. The complex which has facial and meridional isomers is
(Given: py = pyridine and en = $\text{H}_2\text{N}-\text{CH}_2-\text{CH}_2-\text{NH}_2$)

- (1) $[\text{Ni}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}$ (2) $[\text{Cr}(\text{py})_3(\text{Cl})_3]$
(3) $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$ (4) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{3+}$

Ans. (2)

Sol. Ma_3B_3 shows facial and meridional isomerism.

62. The numbers 17.0145 and 21.0235 were rounded to three figures after the decimal point. The resulting numbers, respectively, are

- (1) 17.015 and 21.024 (2) 17.014 and 21.023
(3) 17.015 and 21.023 (4) 17.014 and 21.024

Ans. (4)

Sol. $17.0145 = 17.014$

$21.0235 = 21.024$

During rounding off when 5 digit is dropped and preceding digit is even then no change and one unit is added in preceding odd digit.

63. The amount of carbon dioxide evolved upon complete combustion of 116 g of *n*-butane is
(Given: atomic mass in amu H = 1, C = 12 and O = 16)

- (1) 362 g (2) 352 g (3) 322 g (4) 176 g

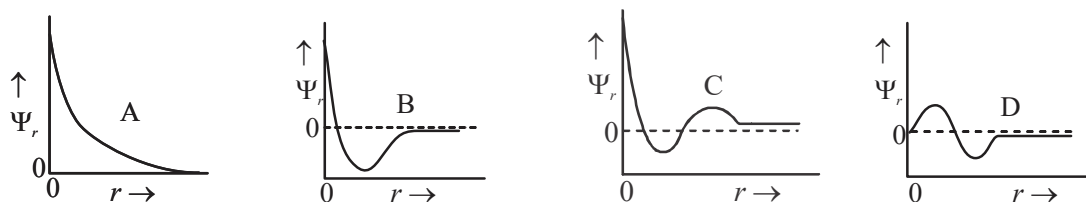
Ans. (2)

Sol. $\text{C}_4\text{H}_{10} + \frac{13}{2}\text{O}_2 \rightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$

58g $4 \times 44\text{g}$

116 g $\frac{4 \times 44 \times 116}{58} = 352\text{g}$

64. Consider the following schematic plots of orbital wavefunction (ψ_r) against distance (r) from nucleus.



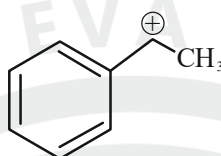
The figure representing two radial nodes in the orbital is

- (1) D (2) A (3) B (4) C

Ans. (4)

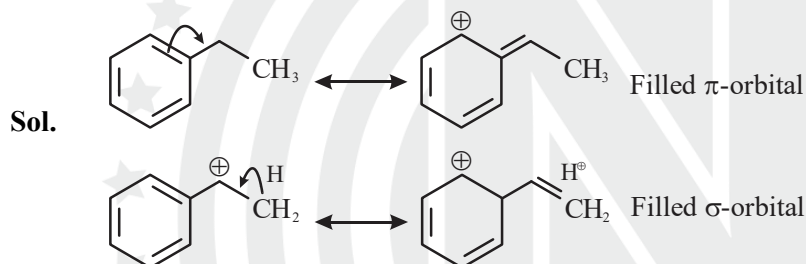
Sol. \therefore Radial node = 2 in graph C.

65. The following carbocation is stabilized by the interaction of the empty p orbital with



- (1) empty σ^* and empty π^* orbitals (2) filled σ and filled π orbitals
 (3) empty σ and empty π^* orbitals (4) empty σ^* and filled π orbitals

Ans. (2)



66. A 1 : 3 electrolyte in an aqueous solution is

- (1) $[\text{Co}(\text{NH}_3)_3(\text{NO}_2)_3]$ (2) $[\text{CoCl}_2(\text{NH}_3)_4]\text{Cl}$
 (3) $[\text{CoCl}(\text{NH}_3)_5]\text{Cl}_2$ (4) $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$

Ans. (4)

Sol. $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$

67. The standard electrode potential (E°) for the half-cell reaction $\text{Fe}^{3+} + e^- \rightarrow \text{Fe}^{2+}$ at 298 K is
 (Given: $E^\circ(\text{Fe}^{3+}/\text{Fe}) = -0.04\text{ V}$ and $E^\circ(\text{Fe}^{2+}/\text{Fe}) = -0.44\text{ V}$ at 298 K)

- (1) +0.92 V (2) +0.40 V (3) +0.76 V (4) -0.48 V

Ans. (3)

Sol. (i) $\text{Fe}^{3+} + 3e^- \rightarrow \text{Fe}$ $\Delta G_1^0 = -3F \times -0.04 = 0.12F$

(ii) $\text{Fe}^{2+} + 2e^- \rightarrow \text{Fe}$ $\Delta G_2^0 = -2F \times -0.44 = 0.88F$

(i) — (ii)

$\text{Fe}^{3+} + 2e^- \rightarrow \text{Fe}^{2+}$ $\Delta G_3^0 = \Delta G_1^0 - \Delta G_2^0 = 0.12F - 0.88F$
 $= -0.76F$

Now, $E^\circ \Rightarrow -1 \times F \times E_{\text{cell}}^\circ = -0.76F$

$E_0 = 0.76\text{ V}$

68. In potash alum, the ratio of K^+ and SO_4^{2-} ions is
 (1) 3 : 2 (2) 1 : 2 (3) 2 : 1 (4) 2 : 3

Ans. (2)

Sol. $2K^+ : 4SO_4^{2-}$
 1 : 2

69. Consider the following statements about the solutions formed by mixing two liquids.
 A. An ideal solution thus formed obeys Raoult's law throughout the composition range.
 B. Mixture of chloroform and acetone shows negative deviation from Raoult's law.
 C. Mixture of aniline and phenol shows positive deviation from Raoult's law.
 (1) A and C only (2) A and B only (3) B and C only (4) A only

Ans. (2)

Sol. A and B Correct C is incorrect. (Factual)

70. For a salt XY, which is a strong electrolyte, the plot of Λ_m versus \sqrt{c} has a slope of $-90.0 \text{ S cm}^2 \text{ mol}^{-3/2} \text{ L}^{1/2}$ at 298 K. At 0.01 M concentration of XY, the value of Λ_m is $145.0 \text{ S cm}^2 \text{ mol}^{-1}$. The limiting molar conductivity of Y^- ion ($\lambda_{Y^-}^0$, in $\text{S cm}^2 \text{ mol}^{-1}$) at 298 K will be

(Given: $\lambda_{X^+}^0 = 74.0 \text{ S cm}^2 \text{ mol}^{-1}$)

- (1) 76.0 (2) 80.0 (3) 100.0 (4) 90.0

Ans. (2)

Sol. Deby-Huckel concentration equation—

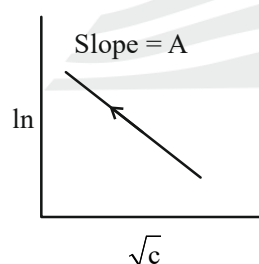
$$\Lambda_m = \Lambda_m^0 - A\sqrt{C}$$

$$14S = \Lambda_m^0 - 90 \times \sqrt{0.01}$$

$$\Lambda_m^0 = 14S + 9.0 = 154 \text{ S cm}^2$$

$$\text{Now, } \Lambda_{xy}^0 = \Lambda_x^0 + \Lambda_y^0$$

$$154 = 74 + \Lambda_y^0 \Rightarrow \Lambda_y^0 = 154 - 74 = 80 \text{ S cm}^2 / \text{mol}$$



71. Arrange the following compounds in the increasing order of polarity
 A. $CH_3CH_2OCH_2CH_3$ B. CH_3CH_2OH
 C. CH_3COCH_3 D. CH_3COOH

Choose the correct answer from the options given below.

- (1) $A < C < B < D$ (2) $A < B < C < D$
 (3) $C < A < D < B$ (4) $C < A < B < D$

Ans. (1)

Sol. $A < C < B < D$ (Factual)

72. According to crystal field theory, the correct order of ligands with respect to their decreasing order of field strength is

- (1) $\text{Cl}^- > \text{NH}_3 > \text{H}_2\text{O} > \text{CO}$ (2) $\text{CO} > \text{NH}_3 > \text{H}_2\text{O} > \text{Cl}^-$
(3) $\text{CO} > \text{H}_2\text{O} > \text{NH}_3 > \text{Cl}^-$ (4) $\text{Cl}^- > \text{H}_2\text{O} > \text{NH}_3 > \text{CO}$

Ans. (2)

Sol. $\text{CO} > \text{NH}_3 > \text{H}_2\text{O} > \text{Cl}^-$ (Factual)

73. The amino acid that gives a red-blood colour on treating its sodium fusion extract with sodium nitroprusside is

- (1) serine (2) leucine (3) threonine (4) methionine

Ans. (4)

Sol. Sulphur containing amino acid gives blood colour in nitroprusside tests.

74. In an acidic medium, 10 mL of 0.25 M oxalic acid is titrated with KMnO_4 solution. If the volume of KMnO_4 solution required to reach end point is 10 mL, the strength of the KMnO_4 solution is

- (1) 0.15 M (2) 0.10 M (3) 0.20 M (4) 0.25 M

Ans. (2)

Sol. $N_1 V_1 = N_2 V_2$

Normality oxalic acid = $0.25 \times 2 = 0.5$

Normality $\text{KMnO}_4 = ?$

$$0.5 \times 10 = 10 \times N_2$$

$$N_2 = 5/10 = 0.5 \text{ N}$$

$$\text{Molarity} = \frac{N}{\text{n factor}} = \frac{0.5}{5} = 0.1 \text{ M}$$

75. The correct statement is

- (1) Aluminium has five valence orbitals.
(2) Boron has a maximum covalency of four.
(3) Beryllium has three valence orbitals.
(4) Magnesium has a maximum covalency of four.

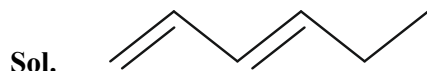
Ans. (2)

Sol. Boron has a max. covalency of four. (Factual)

76. Among the following, the compound having conjugated double bonds is

- (1) hepta-1,6-diene (2) hepta-1,3-diene (3) hepta-1,4-diene (4) hepta-1,5-diene

Ans. (2)



77. $2A \xrightarrow{k} B$ is a zero-order reaction, where $k = 1.0 \text{ mol L}^{-1} \text{ min}^{-1}$. If the initial concentration of A is 2 M, then the time taken to complete 75% of the reaction will be

- (1) 2.0 min (2) 1.5 min (3) 0.75 min (4) 1.0 min

Ans. (3)

Sol. $2A \rightarrow B \Rightarrow \frac{1}{2} \left(\frac{dA}{dt} \right) = K$

$$\frac{-d[A]}{dt} = 2K$$

$$(A)_t = [A]_0 - 2Kt$$

$$2t = 2 - 0.5 = 1.5$$

$$t = 0.75 \text{ min}$$

78. The correct order of solubility of the given salts in water at 298 K is

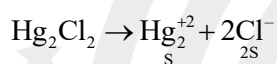
| Salt | K_{sp} at 298 K |
|---------------------------------|-----------------------|
| AgBr | 5.0×10^{-13} |
| Zn(OH) ₂ | 1.0×10^{-15} |
| Hg ₂ Cl ₂ | 1.3×10^{-18} |

- (1) Zn(OH)₂ > AgBr > Hg₂Cl₂ (2) Hg₂Cl₂ > Zn(OH)₂ > AgBr
 (3) AgBr > Zn(OH)₂ > Hg₂Cl₂ (4) Hg₂Cl₂ > AgBr > Zn(OH)₂

Ans. (1)

Sol. AgBr - $K_{sp} = S^2 \Rightarrow S = \sqrt{K_{sp}} = \sqrt{5 \times 10^{-13}}$
 $= 7.07 \times 10^{-7}$

Zn(OH)₂, $K_{sp} = 4s^3 \Rightarrow S = 3\sqrt{K_{sp}/4} = \sqrt[3]{\frac{1.0 \times 10^{-15}}{4}} = 6.3 \times 10^{-6}$



$$4S^3 = K_{sp}$$

$$S = 3\sqrt{K_{sp}/4} = 3\sqrt{\frac{1.3 \times 10^{-18}}{4}} = 6.87 \times 10^{-7}$$



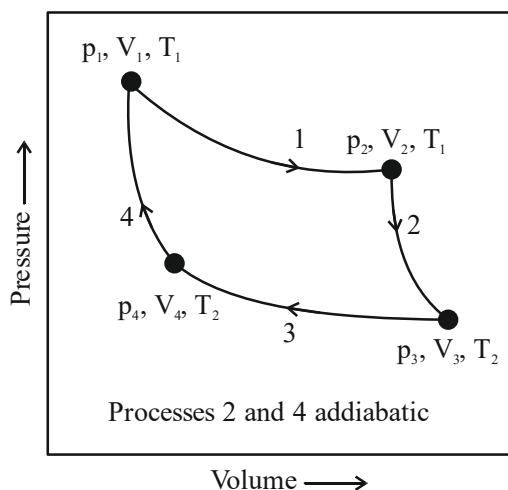
79. The correct decreasing order of oxidation state of the underlined atom in each molecule is

- (1) P₄O₆ > Cl₂O₇ > AlH₃ (2) P₄O₁₀ > SO₃ > H₂O
 (3) N₂O₅ > Al₂O₃ > H₂S (4) PbO₂ > N₂O₃ > SO₃

Ans. (3)

Sol. N_2O_3 > Al_2O_3 > H_2S
₊₅ ₊₃ ₋₂

80. Consider the reversible processes for 1.0 mol of an ideal gas as shown in the figure.



$$\Delta H^\circ = T\Delta S^\circ$$

$$\Delta S^\circ = \frac{\Delta H^\circ}{T} = \frac{666}{333} = 2 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

84. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason R.

Assertion A: Generally, 3d transition metals have high melting points.

Reason R: Involvement of 3d-electrons in addition to 4s-electrons in the interatomic metallic bonding.

In light of the above statements, choose the most appropriate answer from the options given below:

- (1) A is not correct but R is correct.
- (2) Both A and R are correct and R is the correct explanation of A.
- (3) Both A and R are correct and R is NOT the correct explanation of A.
- (4) A is correct but R is not correct.

Ans. (2)

Sol. In Transition metal energy of 3d orbital is very close to 4s orbital as a result both e^- of 3d and 4s involve in metallic bonding.

85. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason R.

Assertion A: The first ionization enthalpy of O is lower than that of N and F.

Reason R: The loss of an electron from O leads to stable half-filled p orbital.

In light of the above statements, choose the most appropriate answer from the options given below:

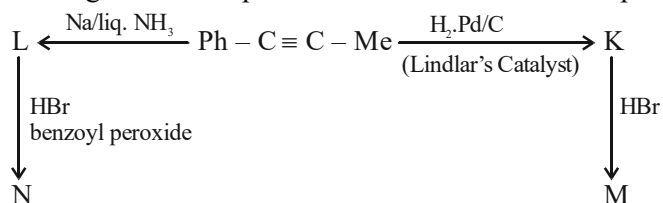
- (1) A is not correct but R is correct.
- (2) Both A and R are correct and R is the correct explanation of A.
- (3) Both A and R are correct and R is NOT the correct explanation of A.
- (4) A is correct but R is not correct.

Ans. (2)

Sol. I.E.: $O < N < F$

After removal of e^- from O, O^+ becomes $2p^3$ (half-filled stable)

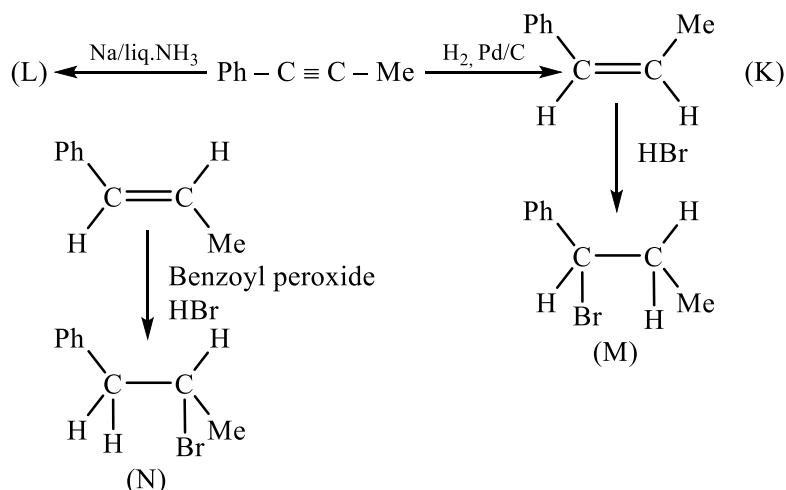
86. Consider the following reaction sequences and choose the correct option.



- (1) M and N stereoisomers
- (2) K and L are geometrical isomers
- (3) K and L are enantiomers
- (4) M and N are geometrical isomers

Ans. (2)

Sol.



K and L are G.I

87. The highest occupied molecular orbital for Ne_2 is

- (1) σ_{2p}^* (2) π_{2p} (3) σ_{2p} (4) π_{2p}^*

Ans. (1)

Sol. σ_{2p}^* : H.O.M.O

88. Match the species in List-I with their geometry in List-II.

List-I

A. PCl_5

B. BrF_5

C. BF_4^-

D. $[\text{Ni}(\text{CN})_4]^{2-}$

List-II

I. Tetrahedral

II. Square Planar

III. Trigonal bipyramidal

IV. Square pyramidal

Choose the correct answer from the options given below:

(1) A—III, B—II, C—I, D—IV

(2) A—IV B—III, C—I, D—II

(3) A—III, B—IV, C—I, D—II

(4) A—III, B—I, C—II, D—IV

Ans. (3)

Sol. PCl_5 : sp^3d

BrF_5 : sp^3d^2

BF_4^- : sp^3

$[\text{Ni}(\text{CN})_4]^{2-}$: dsp^2

89. Match the vitamin in List-I with their source in List-II.

List-I

A. vitamin A

B. vitamin B_{12}

C. vitamin E

D. vitamin K

List-II

I. meat

II. Sunflower oil

III. Green leafy vegetables

IV. carrots

Choose the correct answer from the options given below:

(1) A—III, B—I, C—IV, D—II

(2) A—II B—III, C—IV, D—I

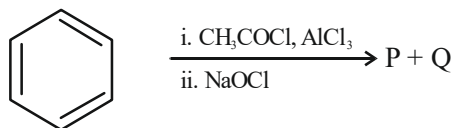
(3) A—IV, B—I, C—II, D—III

(4) A—IV, B—II, C—I, D—III

Ans. (3)

Sol. Fact

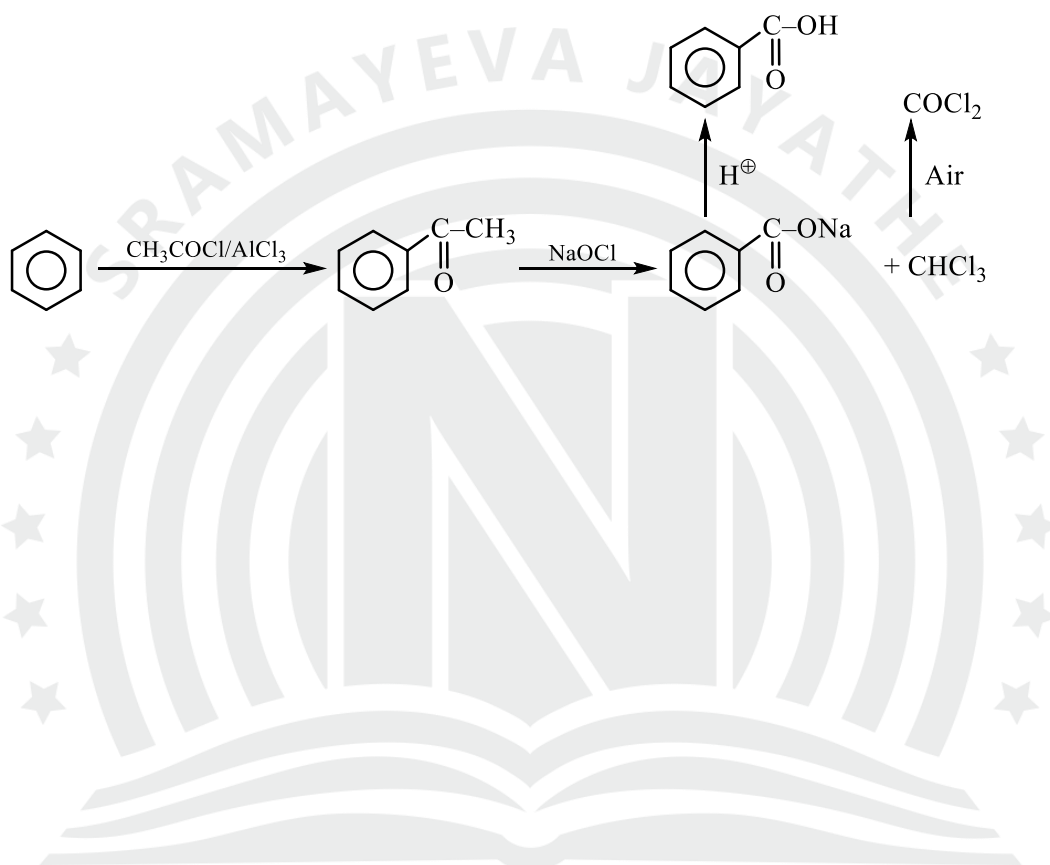
90. For the following reaction sequence, choose the correct option



- (1) Both **P** and **Q** are carbonyl compounds.
- (2) If **P** is the sodium salt of a carboxylic acid, **Q** is a primary alcohol.
- (3) **P** and **Q** are aromatic compounds.
- (4) If **P** gives a carboxylic acid on acidification, **Q** gives a poisonous gas on exposure to air and light.

Ans. (4)

Sol.



BIOLOGY

91. Given below are two statements

Statement I : The class name Reptilia refers to creeping or crawling mode of locomotion.

Statement II : All organisms belonging to Reptilia have three chambered heart.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Statement I is incorrect but Statement II is correct
- (2) Both Statement I and Statement II are correct
- (3) Both Statement I and Statement II are incorrect
- (4) Statement I is correct but Statement II is incorrect

Ans. 4

92. How many turns of Calvin cycle are required for the formation of three molecules of glucose ?

- (1) 18
- (2) 6
- (3) 3
- (4) 1

Ans. 1

93. Photorespiration reaction catalyzed by RuBisCo is shown below :



Identify " X " from the given options:

- (1) Malate
- (2) Phosphoenolpyruvate
- (3) 2-Phosphoglycolate
- (4) Oxaloacetate

Ans. 3

94. Given below are two statements :

Statement I : In gymnosperms, the male and female gametophytes remain within the sporangia.

Statement II : In gymnosperms, seeds are not covered.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Statement I is incorrect but Statement II is correct
- (2) Both Statement I and Statement II are correct
- (3) Both Statement I and Statement II are incorrect
- (4) Statement I is correct but Statement II is incorrect

Ans. 2

95. How many molecules of pyruvic acid are produced at the end of glycolysis from 206 molecules of glucose ?

- (1) 412
- (2) 206
- (3) 309
- (4) 103

Ans. 1

96. Match List-I with List-II.

List-I

- A. Fusion of protoplasts between gametes
- B. Fusion of two nuclei
- C. Generation of haploid spores

List-II

- I. Meiosis
- II. Plasmogamy
- III. Karyogamy

Choose the correct answer from the options given below :

- (1) A-I, B-III, C-II
- (2) A-II, B-III, C-I
- (3) A-II, B-I, C-III
- (4) A-III, B-II, C-I

Ans. 2

97. Mitochondrial inner membrane encloses _____.
 (1) aqueous humor (2) matrix (3) cytosol (4) mucus
Ans. 2
98. Phyllotaxy is the pattern of arrangement of _____.
 (1) sepals (2) leaves (3) flowers (4) fruits
Ans. 2
99. Mad cow disease is caused by _____.
 (1) *Mycoplasma sp.* (2) prions
 (3) viroids (4) *Aspergillus sp.*
Ans. 2
100. Cell theory was formulated by _____.
 (1) Antonie Von Leeuwenhoek
 (2) Schleiden and Schwann
 (3) Robert Brown
 (4) Singer and Nicolson
Ans. 2
101. Which of the following plant growth regulators promotes internode elongation prior to flowering in cabbage?
 (1) Ethephon (2) Abscisic acid
 (3) Gibberellin (4) Indole butyric acid
Ans. 3
102. Which pigment has absorption peak at 700 nm in the photosynthetic reaction centre PS I (P700)?
 (1) Carotenoids (2) Chlorophyll b (3) Chlorophyll a (4) Xanthophylls
Ans. 3
103. Sphenopsida class belongs to _____.
 (1) pteridophytes (2) bryophytes (3) angiosperms (4) gymnosperms
Ans. 1
104. Which of the following represents the correct sequence of arrangement of bones in the lower limb of humans ?
 (1) Femur-tarsal-patella-tibia (2) Femur-tibia-patella-tarsal
 (3) Patella-femur-tibia-tarsal (4) Femur-patella-tibia-tarsal
Ans. 4
105. Which of the following plant growth regulators is used as herbicide ?
 (1) Gibberellin (2) 2,4-D (3) Kinetin (4) Abscisic acid
Ans. 2
106. Genus represents _____.
 (1) a group of closely related families (2) an individual plant or animal
 (3) a population of plants and animals (4) a group of closely related species
Ans. 4
107. The plastid that stores xanthophyll is known as _____.
 (1) amyloplast (2) chloroplast (3) chromoplast (4) aleuroplast
Ans. 3
108. In water, frogs respire using _____.
 (1) trachea (2) skin (3) buccal cavity (4) lungs
Ans. 2

- 109.** Which of the following is not a characteristic of chordates?
(1) Presence of post anal part (tail) (2) Presence of notochord
(3) Central nervous system is dorsal (4) Absence of gills

Ans. 4

- 110.** Smooth endoplasmic reticulum _____.
(1) is a site for the synthesis of carbohydrates
(2) has ribosomes attached to its surface
(3) is the major site for the synthesis of lipids
(4) is actively involved in protein synthesis

Ans. 3

- 111.** Which of the following are characteristics of prokaryotic cells?

- (a) Ribosomes are made of 50 S and 30 S subunits
(b) They can have plasmids
(c) They contain mesosome
(d) They have peroxisomes

Choose the correct answer from the options given below:

- (1) (a), (b) and (c) only (2) (b) and (c) only
(3) (a) and (c) only (4) (a), (c) and (d) only

Ans. 1

- 112.** Match List-I with List-II.

- | List-I | List-II |
|-----------------|------------------------------------------------|
| A. Cristae | I. Flat membrane sacs in stroma of chloroplast |
| B. Cisternae | II. Infoldings in mitochondria |
| C. Thylakoids | III. Cell membrane |
| D. Phospholipid | IV. Disc shaped sacs in the Golgi apparatus |

Choose the correct answer from the options given below:

- (1) A-IV, B-III, C-I, D-II (2) A-III, B-IV, C-I, D-II
(3) A-II, B-IV, C-I, D-III (4) A-II, B-IV, C-III, D-I

Ans. 3

- 113.** Which of the following statements related to pituitary gland are correct

- (a) It is divided anatomically into adenohypophysis and neurohypophysis
(b) It secretes follicle stimulating hormone
(c) It secretes melanocyte stimulating hormone
(d) It does not secrete prolactin

Choose the correct answer from the options given below:

- (1) (b) and (c) only (2) (a) and (b) only
(3) (a), (b) and (c) only (4) (c) and (d) only

Ans. 3

- 114.** Which of the following statements regarding photorespiration are correct?

- (a) Do not occur in C₃ plants
(b) CO₂ is consumed and O₂ is generated
(c) Phosphoglycolate is formed
(d) No synthesis of ATP and NADPH

Choose the correct answer from the options given below:

- (1) (a) and (b) only (2) (a) and (d) only (3) (c) and (d) only (4) (b) and (d) only

Ans. 3

115. Which of the following statements is incorrect?

- (1) Fibrinogen is produced from fibrin
- (2) Blood coagulates in response to an injury
- (3) Blood clot consists of fibrins
- (4) Fibrin is produced from fibrinogen

Ans. 1

116. Arrange the following taxonomic categories in ascending order.

- (a) Genus (b) Class (c) Order (d) Phylum
(e) Family (f) Kingdom (g) Species

Choose the correct answer from the options given below :

- (1) (f), (c), (b), (g), (d), (e), (a)
- (2) (g), (a), (e), (c), (b), (d), (f)
- (3) (a), (c), (d), (g), (f), (b), (e)
- (4) (g), (c), (d), (b), (e), (a), (f)

Ans. 2

117. Select the correct sequence of experiments that led to a gradual understanding of photosynthesis in green plants.

- (1) Production of glucose → role of air → release of oxygen → absorption spectra of chlorophyll a and b
- (2) Absorption spectra of chlorophyll a and b → production of glucose → release of oxygen → role of air
- (3) Role of air → release of oxygen → production of glucose → absorption spectra of chlorophyll a and b
- (4) Release of oxygen → production of glucose → absorption spectra of chlorophyll a and b → role of air

Ans. 3

118. Match List-I with List-II.

- | List-I | List-II |
|-------------------|------------------------|
| A. Starch | I. Fights infection |
| B. Antibody | II. Energy storage |
| C. Concanavalin A | III. Glucose transport |
| D. Glut-4 | IV. Lectin |

Choose the correct answer from the options given below :

- (1) A-I, B-II, C-III, D-IV
- (2) A-I, B-II, C-IV, D-III
- (3) A-II, B-I, C-IV, D-III
- (4) A-II, B-I, C-III, D-IV

Ans. 3

119. The number of vertebrae in a human is _____.

- (1) 206 (2) 7 (3) 12 (4) 26

Ans. 4

120. Endomembrane system includes _____.

- (1) Golgi complex, chloroplast, peroxisomes and vacuole
- (2) endoplasmic reticulum, Golgi complex, lysosomes and vacuole
- (3) endoplasmic reticulum, chloroplast, peroxisomes and vacuole
- (4) mitochondria, chloroplast, peroxisomes and vacuole

Ans. 2

121. Length of the stem at time 0 is 20 cm. The arithmetic growth rate is 30 cm per day. What is the length of the stem at the end of the 7th day?

- (1) 460 cm (2) 50 cm (3) 170 cm (4) 230 cm

Ans. 4

122. Match List-I with List-II.

| List I | List II |
|--------------|---------------|
| A. Spherical | I. Vibrio |
| B. Rod | II. Cocci |
| C. Comma | III. Spirilla |
| D. Spirillum | IV. Bacilli |

Choose the correct answer from the options given below :

- (1) A-II, B-IV, C-I, D-III (2) A-I, B-III, C-II, D-IV
(3) A-III, B-II, C-I, D-IV (4) A-II, B-I, C-IV, D-III

Ans. 1

123. The number of action potentials generated by sino-arterial node (SAN) in a healthy human is _____ per minute.

- (1) 120-140 (2) 28-30 (3) 70-75 (4) 100-110

Ans. 3

124. Match List-I with List-II.

| List-I | List-II |
|-----------|---------------------|
| A. Family | I. Sapindales |
| B. Genus | II. Dicotyledonae |
| C. Class | III. Anacardiaceae |
| D. Phylum | IV. Angiospermae |
| E. Order | V. <i>Mangifera</i> |

Choose the correct answer from the options given below :

- (1) A-III, B-V, C-II, D-IV, E-I (2) A-I, B-V, C-II, D-IV, E-III
(3) A-II, B-I, C-III, D-IV, E-V (4) A-II, B-III, C-V, D-I, E-IV

Ans. 1

125. Which of the following is not a part of human central neural system ?

- (1) Pericardium (2) Arachnoid (3) Dura mater (4) Pia mater

Ans. 1

126. Given below are two statements :

Statement I : Chromosomes are fully condensed at the end of prophase I.

Statement II : Meiosis I resembles mitosis.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Statement I is incorrect, but Statement II is true
(2) Both Statement I and Statement II are true
(3) Both Statement I and Statement II are false
(4) Statement I is correct, but Statement II is false

Ans. 4

127. Match List-I with List-II.

| List-I | List-II |
|------------------------------|----------------------|
| A. Marginal placentation | I. <i>Argemone</i> |
| B. Axile placentation | II. Tomato |
| C. Parietal placentation | III. <i>Primrose</i> |
| D. Free central placentation | IV. Pea |

Choose the correct answer from the options given below :

- (1) A-IV, B-II, C-I, D-III (2) A-II, B-IV, C-I, D-III
(3) A-IV, B-II, C-III, D-I (4) A-IV, B-III, C-I, D-II

Ans. 1

- 128.** Symbiotic association between fungi and algae are called _____.
 (1) chrysophytes (2) lichens (3) sponges (4) mycorrhiza
Ans. 2
- 129.** Which of the following is not a prokaryote ?
 (1) Fungi (2) Bacteria (3) Blue green algae (4) Mycoplasma
Ans. 1
- 130.** Arrange the following elements in descending order of their contribution to percentage weight of the human body.
 (a) Oxygen (b) Carbon (c) Hydrogen (d) Nitrogen
 Choose the correct answer from the options given below :
 (1) (b), (a), (c), (d) (2) (a), (b), (c), (d) (3) (c), (a), (b), (d) (4) (b), (c), (d), (a)
Ans. 2
- 131.** Which one of the following statements is incorrect?
 (1) β -cells of pancreas secrete insulin (2) α -cells of pancreas secrete glucagon
 (3) α -cells of pancreas secrete insulin (4) Glucagon stimulates glycogenolysis
Ans. 3
- 132.** Which of the following are characteristic features of Solanaceae family?
 (a) Flowers are bisexual and actinomorphic
 (b) Calyx have five sepals and are united
 (c) Androecium have five stamens and are epipetalous
 (d) Ovary is inferior
 Choose the correct answer from the options given below :
 (1) (b), (c) and (d) only (2) (a), (b) and (c) only
 (3) (d) only (4) (a) and (b) only
Ans. 2
- 133.** Given below are two statements :
Statement I : When any plane passing through the central axis of the body divides the organism into two identical, halves, it is called radial symmetry.
Statement II : In phylum Echinodermata, both adults and larvae are radially symmetrical.
 In the light of the above statements, choose the most appropriate answer from the options given below:
 (1) Statement I is incorrect but Statement II is correct
 (2) Both Statement I and Statement II are correct
 (3) Both Statement I and Statement II are incorrect
 (4) Statement I is correct but Statement II is incorrect
Ans. 4
- 134.** The correct sequence of adult cell cycle phases is _____.
 (1) S-M-G₂-G₁ (2) G₁-G₂-S-M (3) G₁-M-G₂-S (4) G₁-S-G₂-M
Ans. 4
- 135.** In frogs, the number of pairs of cranial nerves arising from the brain are _____.
 (1) 12 (2) 6 (3) 9 (4) 10
Ans. 4

136. Given below are two statements : one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A: In recombinant DNA technology, lysozyme is used for disrupting bacterial cells while cellulase is for plant cells.

Reason R: Isolation of genetic material needs disruption of cells.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) A is not correct but R is correct
- (2) Both A and R are correct and R is the correct explanation of A
- (3) Both A and R are correct but R is not the correct explanation of A
- (4) A is correct but R is not correct

Ans. 2

137. The method of directly of injecting a sperm into ovum in assisted reproductive technology is called :

- (1) Embryo transfer (ET)
- (2) Gamete intra fallopian transfer (GIFT)
- (3) Zygote intra fallopian transfer (ZIFT)
- (4) Intra cytoplasmic sperm injection (ICSI)

Ans. 4

138. Adaptive radiation in placental mammals and Australian Marsupials leading to similarity between distant species is an example of ____.

- (1) genetic drift
- (2) divergent evolution
- (3) convergent evolution
- (4) founder effect

Ans. 3

139. Given below are two statements : one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : In an experiment, Mendel observed that the F1 progeny plants are all tall and none are dwarf.

Reason R : Stem height is a contrasting trait, with tall being dominant and dwarf being recessive.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) A is not correct but R is correct
- (2) Both A and R are correct and R is the correct explanation of A
- (3) Both A and R are correct but R is not the correct explanation of A
- (4) A is correct but R is not correct

Ans. 2

140. Arrange the following in descending order of number of species in the Amazonian rain forest.

- (a) Plants
- (b) Birds
- (c) Fishes
- (d) Invertebrates
- (e) Mammals

Choose the correct answer from the options given below:

- (1) (b) > (a) > (d) > (c) > (e)
- (2) (c) > (b) > (d) > (e) > (a)
- (3) (d) > (a) > (c) > (b) > (e)
- (4) (e) > (b) > (a) > (c) > (d)

Ans. 3

141. Sponges exchange O₂ with CO₂ by _____.

- (1) gills
- (2) simple diffusion over their entire body surfaces
- (3) moist cuticle
- (4) tracheal tubes

Ans. 2

150. Given below are two statements :

Statement I: Ovulation is caused by LH surge leading to rupture of Graafian follicles.

Statement II : Graafian follicle remaining after ovulation transform into corpus luteum and secretes large amount of estrogen.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Statement I is incorrect but Statement II is correct
- (2) Both Statement I and Statement II are correct
- (3) Both Statement I and Statement II are incorrect
- (4) Statement I is correct but Statement II is incorrect

Ans. 4

151. The opening between the right atrium and the right ventricle is guarded by _____.

- (1) sino-atrial node
- (2) bicuspid valve
- (3) tricuspid valve
- (4) semilunar valve

Ans. 3

152. Which of the following is not evidence for evolution?

- (1) Divergent evolution of anatomical structures such as forelimbs
- (2) Convergent evolution of traits like wings of birds and butterflies
- (3) Paleontological evidence from fossil records
- (4) Embryological support for evolution as proposed by Ernst Heckel

Ans. 4

153. The inactive form of Bt toxin is converted to the active form in the insect gut _____.

- (1) by nucleases
- (2) due to alkaline pH
- (3) due to acidic pH
- (4) by proteases

Ans. 2

154. Colostrum, secreted by mother during initial days of lactation, is abundant in _____.

- (1) IgD
- (2) IgG
- (3) IgM
- (4) IgA

Ans. 4

155. Which of the following in female gametophyte of an angiosperm helps in guiding the pollen tube for fertilizing the eggs?

- (1) Polar nucleus
- (2) Antipodals
- (3) Synergids
- (4) Central cells

Ans. 3

156. Which of the following disease is not sexually transmitted?

- (1) Genital warts
- (2) Syphilis
- (3) Tuberculosis
- (4) Gonorrhoea

Ans. 3

157. Which of the following statements about lac-operon is correct?

- (1) Galactose can act as an inducer of lac operon
- (2) Gene i is constitutively expressed
- (3) Lactose activates repressor to bind to the operator
- (4) Genes i, z, y and a share single common promoter

Ans. 2

158. Match List-I with List-II.

| List-I | List-II |
|-------------------|-----------------------------------|
| A. Transformation | I. Restriction enzyme |
| B. Cloning site | II. Transfer DNA to host bacteria |
| C. Selection | III. Replication |
| D. Ori | IV. Antibiotic |

Choose the correct answer from the options given below:

- (1) A-IV, B-I, C-III, D-II (2) A-II, B-I, C-IV, D-III
(3) A-I, B-II, C-IV, D-III (4) A-III, B-IV, C-II, D-I

Ans. 2

159. A population of diploid organisms is at Hardy-Weinberg equilibrium. If the frequency of allele A is 0.1, the frequency of AA is _____.

- (1) 0.99 (2) 0.01 (3) 0.02 (4) 0.10

Ans. 2

160. Sperm motility is due to _____.

- (1) muscular movement (2) flagellar movement
(3) ciliary movement (4) amoeboid movement

Ans. 2

161. Consider a population of 10 million cells, Given the per-capita birth rate of 0.002 (per unit time) and the per-capita death rate of 0.002 (per unit time), the expected number of cells after 10 generations is _____.

- (1) 100 million (2) 1 million (3) 5 million (4) 10 million

Ans. 4

162. Given below are two statements : one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : Forelimbs of human and bats are homologous.

Reason R : Forelimbs of humans and bats have similar anatomical structure.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) A is false but R is true
(2) Both A and R are correct and R is the correct explanation of A
(3) Both A and R are true, but R is not the correct explanation of A
(4) A is true but R is false

Ans. 2

163. Muscle contraction is initiated by a signal sent by the central nervous system by the release of _____.

- (1) cyclic adenine monophosphate (2) acetyl choline
(3) acetyl coenzyme A (4) cyclic guanine monophosphate

Ans. 2

164. Which of the following hormone is not secreted by human placenta?

- (1) LH (2) hCG (3) Estrogen (4) Progesterone

Ans. 1

165. Which of the following statements is correct about *Plasmodium*?

- (1) Fertilization takes place in mosquito gut (2) Reproduces sexually in liver cells
(3) Reproduces sexually in RBCs (4) Gametocytes develop in mosquito gut

Ans. 1

171. If the diploid chromosome number of typical angiosperm is 36, what would be the chromosome number in its endosperm?

- (1) 72 (2) 18 (3) 36 (4) 54

Ans. 4

172. Which of the following enzymes synthesizes precursor mRNA?

- (1) DNA polymerase (2) RNA polymerase I
(3) RNA polymerase II (4) RNA polymerase III

Ans. 3

173. Given below are two statements :

Statement I : Plasmids are autonomously replicating DNA.

Statement II : Plasmids are extrachromosomal DNA.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Statement I is incorrect but Statement II is correct
(2) Both Statement I and Statement II are correct
(3) Both Statement I and Statement II are incorrect
(4) Statement I is correct but Statement II is incorrect

Ans. 2

174. How many theca are present in each lobe of a typical bilobed angiosperm anther?

- (1) 12 (2) 2 (3) 6 (4) 8

Ans. 2

175. Natural selection can lead to _____.

- (a) stabilisation (b) genetic drift
(c) directional change (d) disruption

Choose the correct answer from the options given below:

- (1) (a) and (c) only (2) (a) only
(3) (a), (c) and (d) only (4) (a), (b), (c) and (d)

Ans. 3

176. Which of the following statements are correct?

- (a) Energy flow from producers to consumers is unidirectional
(b) Energy pyramid can never be inverted
(c) Transfer of energy follows the 1% law

Choose the correct answer from the options given below:

- (1) (b) and (c) only (2) (a), (b) and (c)
(3) (a) and (b) only (4) (a) and (c) only

Ans. 3

177. Match List-I with List-II.

List-I

- A. Excess growth hormone
B. Luteinizing hormone
C. Vasopressin
D. Oxytocin

List-II

- I. Reabsorption of water and electrolytes in kidney
II. Contraction of uterus during child birth
III. Acromegaly
IV. Ovulation

Choose the correct answer from the options given below:

- (1) A-IV, B-III, C-I, D-II (2) A-III, B-IV, C-II, D-I
(3) A-III, B-IV, C-I, D-II (4) A-II, B-IV, C-I, D-III

Ans. 3

178. Which of the following are secondary lymphoid organs?
(a) Bone marrow (b) Tonsils (c) Spleen (d) Thymus

Choose the correct answer from the option given below:

- (1) (a) and (d) only (2) (a) and (b) only
(3) (b) and (c) only (4) (b) and (d) only

Ans. 3

179. During PCR, primers bind to the DNA strands in the _____ step.

- (1) ligation (2) denaturation (3) extension (4) annealing

Ans. 4

180. Given below are two statements:

Statement I: Down's syndrome is caused by the absence of one of the X-chromosomes.

Statement II: Turner's syndrome is caused by the presence of an additional copy of the chromosomes.

In the light of the above statements, choose the correct answer from the options given below:

- (1) Statement I is incorrect but Statement II is correct
(2) Both Statement I and Statement II are correct
(3) Both Statement I and Statement II are incorrect
(4) Statement I is correct but Statement II is incorrect

Ans. 3



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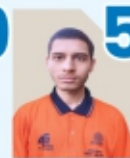
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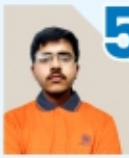
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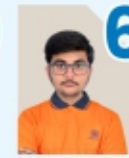
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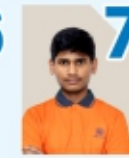
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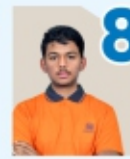
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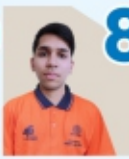
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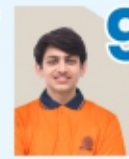
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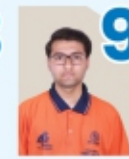
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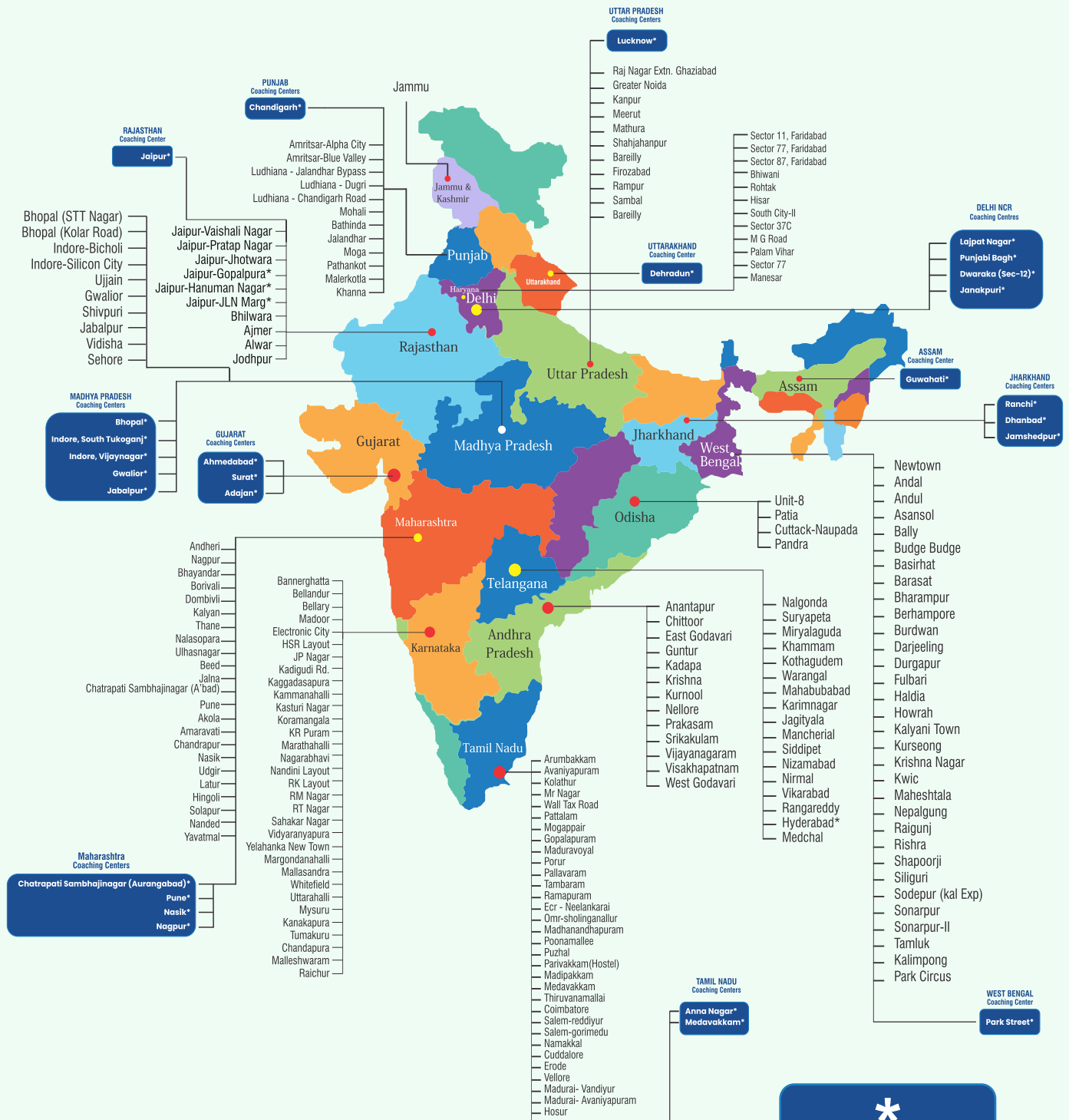
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