

47
YEARS
OF EXCELLENCE



NARAYANA
IIT-JEE/NEET/FOUNDATION

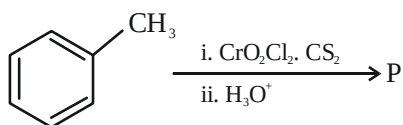
RE-NEET (UG) **2026**



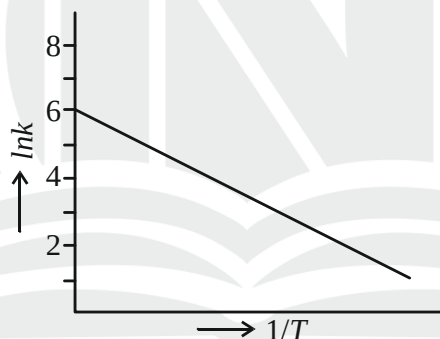
QUESTIONS PAPER
CHEMISTRY
PAPER CODE -70

CHEMISTRY

46. Consider the following reaction, and chose the correct option.



- (1) Compound **P** is obtained by the hydrogenation of benzoyl chloride with Pd on BaSO_4 .
- (2) On treating compound **P** with saturated NaHCO_3 solution, brisk effervescence is observed.
- (3) Compound **P** can be prepared by treating benzene with anhydrous b AlCl_3 and CH_3COCl
- (4) On treatment with bromine water, compound **P** gives a white precipitate.
47. The formula of tetraammineaquachloridocobalt (III) chlorides is
- (1) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}_2$
- (2) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2] \times \text{H}_2\text{O}$
- (3) $[\text{Co}(\text{NH}_3)_4]\text{Cl}_3 \times \text{H}_2\text{O}$
- (4) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}$
48. The lanthanide ion having four unpaired electrons is
[Given: Atomic numbers of Ce = 58, Nd = 60, Tb = 65 and Ho = 67]
- (1) Ho^{3+} (2) Nd^{3+} (3) Ce^{3+} (4) Tb^{3+}
49. For an elementary chemical reaction, the Arrhenius plot is given below.

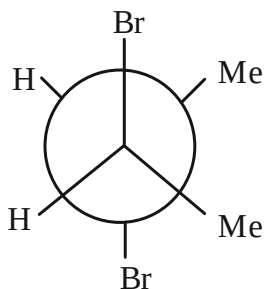


- If the energy of activation is 6.64 kJ mol^{-1} and $R = 8.3 \text{ JK}^{-1} \text{ mol}^{-1}$, the temperature at which the rate constant becomes $\text{e}^2 \text{ min}^{-1}$, is
- (1) 250 K
- (2) 125 K
- (3) 150 K
- (4) 200 K
50. The green paramagnetic species formed by heating KMnO_4 at 513 K is
- (1) KO_2 (2) K_2MnO_4 (3) Mn_3O_4 (4) MnO

51. Given below are two statements:

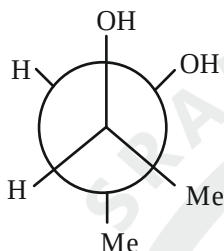
Statement I:

trans-But-2-ene upon treatment with Br_2 in CCl_4 gives the following product



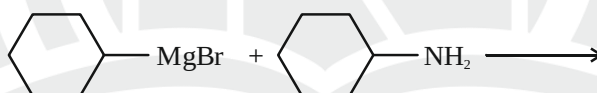
Statement II:

cis-But-2-ene upon treatment with alkaline KMnO_4 gives the following product



In the light of the above statements, choose the **most appropriate** answer from the options given below.

- (1) **Statement I** is incorrect but **Statement II** is correct
(2) Both **Statement I** and **Statement II** are correct
(3) Both **Statement I** and **Statement II** are incorrect
(4) **Statement I** is correct but **Statement II** is incorrect
52. One of the products formed in the following reaction is



- (1)
- (2)
- (3)
- (4)

53. Given below are two statements:

Statement-I: Heating NaCl with concentrated H_2SO_4 and MnO_2 results in oxidation of Mn .

Statement-II: Heating NaI with concentrated H_2SO_4 and MnO_2 results in reduction of Mn .

In light of the above statements, choose the **most appropriate** answer from the options given below:

- (1) Statement-I is incorrect but Statement-II is correct.
(2) Both Statement-I and Statement-II are correct.
(3) Both Statement-I and Statement-II are incorrect.
(4) Statement-I is correct but Statement-II is incorrect.

54. Among the following options, the correct trend in the electron gain enthalpy is

- (1) $I > Br > Cl > F$ (2) $F > Cl > Br > I$
(3) $Br > Cl > F > I$ (4) $Cl > F > Br > I$

55. Given below are two statements:

Statement-I: $[Fe(ox)_3]^{3-}$ is chiral.

Statement-II: $trans-[Cr(H_2O)_2(ox)_2]^-$ is chiral.

In light of the above statements, choose the most appropriate answer from the options given below:

- (1) **Statement-I** is incorrect but **Statement-II** is correct.
(2) Both **Statement-I** and **Statement-II** are correct.
(3) Both **Statement-I** and **Statement-II** are incorrect.
(4) **Statement-I** is correct but **Statement-II** is incorrect.

56. The correct statement about peptides and proteins is

- (1) In α -helices, the polypeptide chain is twisted into a left-handed screw (helix) through intramolecular hydrogen bonds.
(2) Tertiary structure of proteins has two or more polypeptide subunits.
(3) Only the proteins having a quaternary structure are biologically active.
(4) In β -pleated sheet structures, peptide chains are held together by intermolecular hydrogen bonds.

57. Given below are two statements:

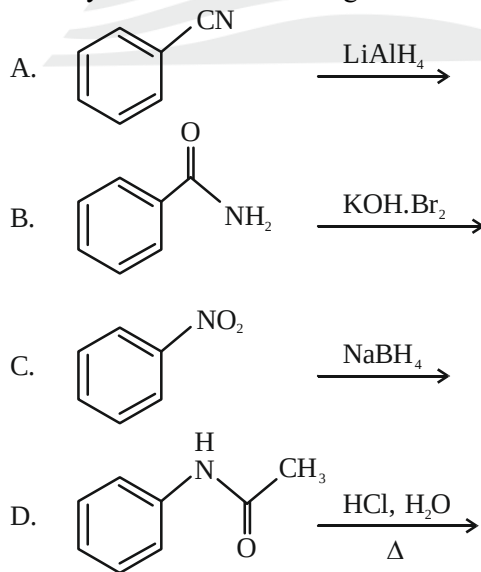
Statement-I: Oxidation of *p* nitrotoluene with acidic $KMnO_4$ gives an acid that is stronger than benzoic acid.

Statement-II: Reduction of *p* nitrotoluene with Sn/HCl followed by neutralization gives an amine that is more basic than aniline.

In light of the above statements, choose the most appropriate answer from the options given below.

- (1) Statement-I is incorrect but Statement-II is correct.
(2) Both Statement-I and Statement-II are correct.
(3) Both Statement-I and Statement-II are incorrect.
(4) Statement-I is correct but Statement-II is incorrect.

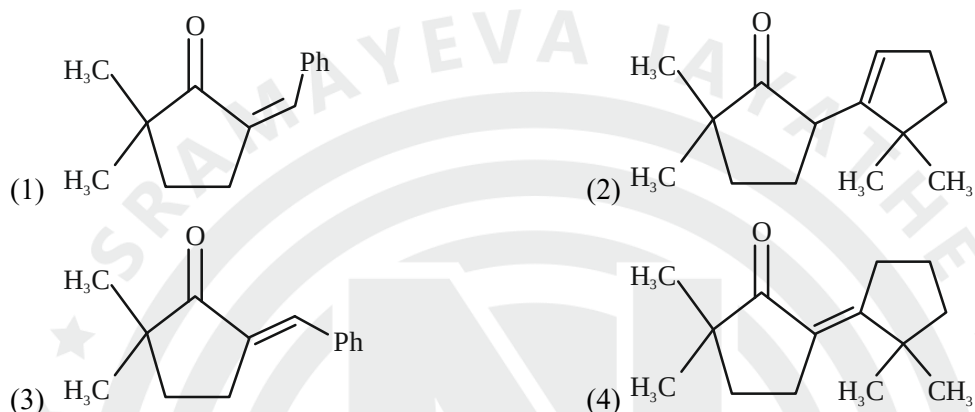
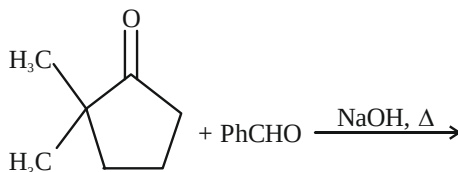
58. Identify the reactions which give aniline as the major product.



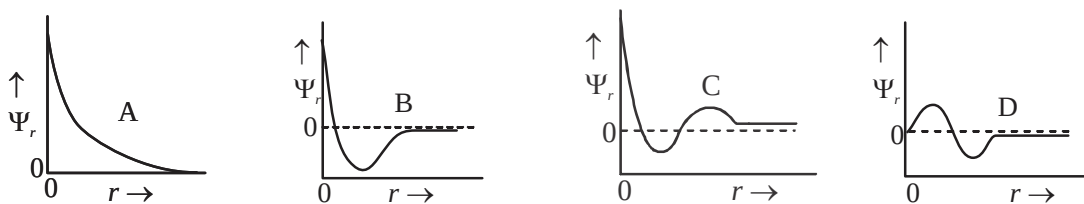
Choose the correct answer from the options given below.

- (1) C and D only (2) A and B only
(3) B and D only (4) A and C only

59. Two moles of an ideal gas undergo free expansion from 10 L to 100 L at 300 K. The values of ΔS_{system} and $\Delta S_{\text{surroundings}}$ are
(R is universal gas constant)
- (1) $\Delta S_{\text{system}} = 4.606 R$; $\Delta S_{\text{surroundings}} = 0$ (2) $\Delta S_{\text{system}} = 0$; $\Delta S_{\text{surroundings}} = 0$
(3) $\Delta S_{\text{system}} = 4.606 R$; $\Delta S_{\text{surroundings}} = -4.606 R$ (4) $\Delta S_{\text{system}} = 0$; $\Delta S_{\text{surroundings}} = 4.606 R$
60. The compound that **CANNOT** be obtained from aldol condensation reaction shown below, is



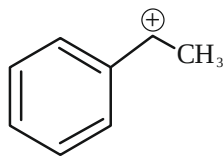
61. The complex which has facial and meridional isomers is
(Given: py = pyridine and en = $\text{H}_2\text{N}-\text{CH}_2-\text{CH}_2-\text{NH}_2$)
- (1) $[\text{Ni}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}$ (2) $[\text{Cr}(\text{py})_3(\text{Cl})_3]$
(3) $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$ (4) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{3+}$
62. The numbers 17.0145 and 21.0235 were rounded to three figures after the decimal point. The resulting numbers, respectively, are
- (1) 17.015 and 21.024 (2) 17.014 and 21.023
(3) 17.015 and 21.023 (4) 17.014 and 21.024
63. The amount of carbon dioxide evolved upon complete combustion of 116 g of *n*-butane is
(Given: atomic mass in amu H = 1, C = 12 and O = 16)
- (1) 362 g (2) 352 g (3) 322 g (4) 176 g
64. Consider the following schematic plots of orbital wavefunction (ψ_r) against distance (r) from nucleus.



The figure representing two radial nodes in the orbital is

- (1) D (2) A (3) B (4) C

65. The following carbocation is stabilized by the interaction of the empty p orbital with

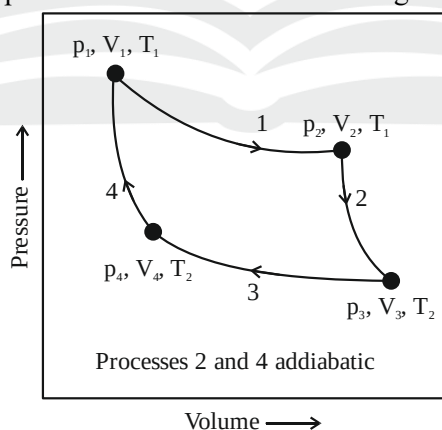


- (1) empty σ^* and empty π^* orbitals (2) filled σ and filled π orbitals
 (3) empty σ and empty π^* orbitals (4) empty σ^* and filled π orbitals
66. A 1 : 3 electrolyte in an aqueous solution is
 (1) $[\text{Co}(\text{NH}_3)_3(\text{NO}_2)_3]$ (2) $[\text{CoCl}_2(\text{NH}_3)_4]\text{Cl}$
 (3) $[\text{CoCl}(\text{NH}_3)_5]\text{Cl}_2$ (4) $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$
67. The standard electrode potential (E°) for the half-cell reaction $\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}^{2+}$ at 298 K is
 (Given: $E^\circ(\text{Fe}^{3+}/\text{Fe}) = -0.04\text{ V}$ and $E^\circ(\text{Fe}^{2+}/\text{Fe}) = -0.44\text{ V}$ at 298 K)
 (1) +0.92 V (2) +0.40 V (3) +0.76 V (4) -0.48 V
68. In potash alum, the ratio of K^+ and SO_4^{2-} ions is
 (1) 3 : 2 (2) 1 : 2 (3) 2 : 1 (4) 2 : 3
69. Consider the following statements about the solutions formed by mixing two liquids.
 A. An ideal solution thus formed obeys Raoult's law throughout the composition range.
 B. Mixture of chloroform and acetone shows negative deviation from Raoult's law.
 C. Mixture of aniline and phenol shows positive deviation from Raoult's law.
 (1) A and C only (2) A and B only (3) B and C only (4) A only
70. For a salt XY, which is a strong electrolyte, the plot of Λ_m versus $\sqrt{c}$ has a slope of $-90.0\text{ S cm}^2\text{ mol}^{-3/2}\text{ L}^{1/2}$ at 298 K. At 0.01 M concentration of XY, the value of Λ_m is $145.0\text{ S cm}^2\text{ mol}^{-1}$. The limiting molar conductivity of Y^- ion ($\lambda_{\text{Y}^-}^0$, in $\text{S cm}^2\text{ mol}^{-1}$) at 298 K will be
 (Given: $\lambda_{\text{X}^+}^0 = 74.0\text{ S cm}^2\text{ mol}^{-1}$)
 (1) 76.0 (2) 80.0 (3) 100.0 (4) 90.0
71. Arrange the following compounds in the increasing order of polarity
 A. $\text{CH}_3\text{CH}_2\text{OCH}_2\text{CH}_3$ B. $\text{CH}_3\text{CH}_2\text{OH}$
 C. CH_3COCH_3 D. CH_3COOH
 Choose the correct answer from the options given below.
 (1) $\text{A} < \text{C} < \text{B} < \text{D}$ (2) $\text{A} < \text{B} < \text{C} < \text{D}$
 (3) $\text{C} < \text{A} < \text{D} < \text{B}$ (4) $\text{C} < \text{A} < \text{B} < \text{D}$
72. According to crystal field theory, the correct order of ligands with respect to their decreasing order of field strength is
 (1) $\text{Cl}^- > \text{NH}_3 > \text{H}_2\text{O} > \text{CO}$ (2) $\text{CO} > \text{NH}_3 > \text{H}_2\text{O} > \text{Cl}^-$
 (3) $\text{CO} > \text{H}_2\text{O} > \text{NH}_3 > \text{Cl}^-$ (4) $\text{Cl}^- > \text{H}_2\text{O} > \text{NH}_3 > \text{CO}$

73. The amino acid that gives a red-blood colour on treating its sodium fusion extract with sodium nitroprusside is
 (1) serine (2) leucine (3) threonine (4) methionine
74. In an acidic medium, 10 mL of 0.25 M oxalic acid is titrated with KMnO_4 solution. If the volume of KMnO_4 solution required to reach end point is 10 mL, the strength of the KMnO_4 solution is
 (1) 0.15 M (2) 0.10 M (3) 0.20 M (4) 0.25 M
75. The correct statement is
 (1) Aluminium has five valence orbitals.
 (2) Boron has a maximum covalency of four.
 (3) Beryllium has three valence orbitals.
 (4) Magnesium has a maximum covalency of four.
76. Among the following, the compound having conjugated double bonds is
 (1) hepta-1,6-diene (2) hepta-1,3-diene (3) hepta-1,4-diene (4) hepta-1,5-diene
77. $2A \xrightarrow{k} B$ is a zero-order reaction, where $k = 1.0 \text{ mol L}^{-1} \text{ min}^{-1}$. If the initial concentration of A is 2 M, then the time taken to complete 75% of the reaction will be
 (1) 2.0 min (2) 1.5 min (3) 0.75 min (4) 1.0 min
78. The correct order of solubility of the given salts in water at 298 K is

Salt	K_{sp} at 298 K
AgBr	5.0×10^{-13}
Zn(OH)_2	1.0×10^{-15}
Hg_2Cl_2	1.3×10^{-18}

- (1) $\text{Zn(OH)}_2 > \text{AgBr} > \text{Hg}_2\text{Cl}_2$ (2) $\text{Hg}_2\text{Cl}_2 > \text{Zn(OH)}_2 > \text{AgBr}$
 (3) $\text{AgBr} > \text{Zn(OH)}_2 > \text{Hg}_2\text{Cl}_2$ (4) $\text{Hg}_2\text{Cl}_2 > \text{AgBr} > \text{Zn(OH)}_2$
79. The correct decreasing order of oxidation state of the underlined atom in each molecule is
 (1) $\underline{\text{P}}_4\text{O}_6 > \underline{\text{Cl}}_2\text{O}_7 > \underline{\text{Al}}\text{H}_3$ (2) $\underline{\text{P}}_4\text{O}_{10} > \underline{\text{S}}\text{O}_3 > \underline{\text{H}}_2\underline{\text{O}}$
 (3) $\underline{\text{N}}_2\text{O}_5 > \underline{\text{Al}}_2\text{O}_3 > \underline{\text{H}}_2\underline{\text{S}}$ (4) $\underline{\text{Pb}}\text{O}_2 > \underline{\text{N}}_2\text{O}_3 > \underline{\text{S}}\text{O}_3$
80. Consider the reversible processes for 1.0 mol of an ideal gas as shown in the figure.



w_1, w_2, w_3 and w_4 represent work done (in calories) in the processes 1, 2, 3 and 4, respectively;
 ΔU_2 and ΔU_4 are changes in the internal energy the processes 2 and 4, respective
 [use $R = 2 \text{ cal K}^{-1} \text{ mol}^{-1}$] The correct option is

- (1) $w_1 + w_2 + w_3 + w_4 = 0$ (2) $w_1 + w_3 = -2T_1 \ln \frac{V_2}{V_1} - 2T_2 \ln \frac{V_4}{V_3}$
 (3) $w_2 + w_4 = \Delta U_2 - \Delta U_4$ (4) $w_1 + w_2 = 2T_1 \ln \frac{V_2}{V_1}$

81. Assertion A:

For an ideal solution formed by mixing liquids **P** and **Q**, $\Delta_{\text{mix}} H = 0$ and $\Delta_{\text{mix}} V = 0$

Reason R:

No interactions occur between **P** and **Q**. In the light of the above statements, choose the most appropriate answer from the options given below.

- (1) **A** is not correct but **R** is correct
- (2) Both **A** and **R** are correct and **R** is the correct explanation of **A**
- (3) Both **A** and **R** are correct but **R** is NOT the correct explanation of **A**
- (4) **A** is correct but **R** is not correct

82. Among the species given below, the spin-only magnetic moment is highest for

(Given: Atomic number of Ti = 22, Mn = 25, Fe = 26 and Co = 27)

- (1) $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ (2) $[\text{Mn}(\text{CN})_6]^{3-}$ (3) $[\text{Fe}(\text{CN})_6]^{3-}$ (4) $[\text{Co}(\text{NH}_3)_6]^{3+}$

83. A protein undergoes reversible thermal denaturation from its initial state **N** to denatured state **D** according to $\text{N} \rightleftharpoons \text{D}$. At 60° , the concentrations of both **N** and **D** are equal at equilibrium, and the standard enthalpy change of denaturation is 666 kJ mol^{-1} . The standard entropy change (ΔS° in $\text{kJ K}^{-1} \text{ mol}^{-1}$) of the protein upon denaturation at 60° closest to

- (1) 11.1 (2) 2.0 (3) 2000.0 (4) 333.0

84. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason **R**.

Assertion A: Generally, 3d transition metals have high melting points.

Reason R: Involvement of 3d-electrons in addition to 4s-electrons in the interatomic metallic bonding.

In light of the above statements, choose the most appropriate answer from the options given below:

- (1) **A** is not correct but **R** is correct.
- (2) Both **A** and **R** are correct and **R** is the correct explanation of **A**.
- (3) Both **A** and **R** are correct and **R** is NOT the correct explanation of **A**.
- (4) **A** is correct but **R** is not correct.

85. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason **R**.

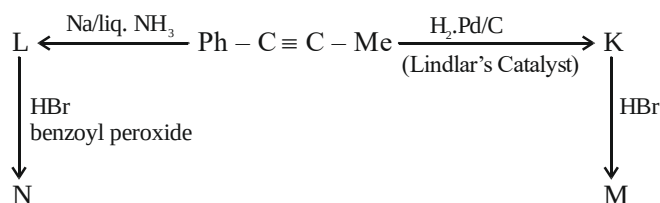
Assertion A: The first ionization enthalpy of O is lower than that of N and F.

Reason R: The loss of an electron from O leads to stable half-filled *p* orbital.

In light of the above statements, choose the **most appropriate** answer from the options given below:

- (1) **A** is not correct but **R** is correct.
- (2) Both **A** and **R** are correct and **R** is the correct explanation of **A**.
- (3) Both **A** and **R** are correct and **R** is NOT the correct explanation of **A**.
- (4) **A** is correct but **R** is not correct.

86. Consider the following reaction sequences and choose the correct option.



- (1) **M** and **N** stereoisomers
 (2) **K** and **L** are geometrical isomers
 (3) **K** and **L** are enantiomers
 (4) **M** and **N** are geometrical isomers
87. The highest occupied molecular orbital for Ne_2 is
 (1) σ_{2p}^* (2) π_{2p} (3) σ_{2p} (4) π_{2p}^*

88. Match the species in List-I with their geometry in List-II.

List-I

A. PCl_5

B. BrF_5

C. BF_4^-

D. $[\text{Ni}(\text{CN})_4]^{2-}$

List-II

I. Tetrahedral

II. Square Planar

III. Trigonal bipyramidal

IV. Square pyramidal

Choose the correct answer from the options given below:

- (1) A—III, B—II, C—I, D—IV
 (2) A—IV B—III, C—I, D—II
 (3) A—III, B—IV, C—I, D—II
 (4) A—III, B—I, C—II, D—IV
89. Match the vitamin in List-I with their source in List-II.

List-I

A. vitamin A

B. vitamin B_{12}

C. vitamin E

D. vitamin K

List-II

I. meat

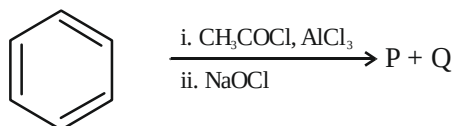
II. Sunflower oil

III. Green leafy vegetables

IV. carrots

Choose the correct answer from the options given below:

- (1) A—III, B—I, C—IV, D—II
 (2) A—II B—III, C—IV, D—I
 (3) A—IV, B—I, C—II, D—III
 (4) A—IV, B—II, C—I, D—III
90. For the following reaction sequence, choose the correct option



- (1) Both **P** and **Q** are carbonyl compounds.
 (2) If **P** is the sodium salt of a carboxylic acid, **Q** is a primary alcohol.
 (3) **P** and **Q** are aromatic compounds.
 (4) If **P** gives a carboxylic acid on acidification, **Q** gives a poisonous gas on exposure to air and light.

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nPrep

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AASHI SINGH
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AIR **12**



SOMYA SHARMA
Classroom Student

AIR **14**



K JEEVAN SAI KUMAR
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AIR **18**



RUPAYAN PAL
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Rijul Jain
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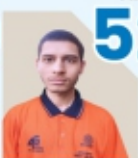
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Naveen Mittal
nPrep

AIR

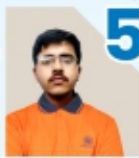
52



Aditya Yadav
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AIR

59



D Surya Charan
Classroom Student

AIR

66



Devyansh Arora
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AIR

70



Y Sameer Kumar
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71



Maulik Bhargamia
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73



Darshan Mule
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75



Yashashwi Kumar
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77



Armaan Bery
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AIR

80



Digant S
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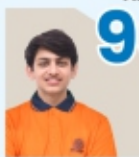
87



Raghav Goyal
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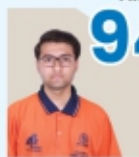
93



Siddh Vimalbhai Vora
Classroom Student

AIR

94



Shah Tejas Twinkle
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AIR

95



Bidisha Majee
Classroom Student

6

Students
in **Top 20**

22

Students
in **Top 100**

94

Students
in **Top 1000**

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TRIUMPHANT SUCCESS

in **NEET (UG)-2024**

MARKS

720

720



1st

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PRACHITA

Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

28th

ARINDAM CHOWDHURY
Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

37th

TANISHK YADAV
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MARKS

715

720



AIR

41st

ADITYA KUMAR PANDA
Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

43rd

ISHA KOTHARI
Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

44th

K SANDEEP CHOWDARY
4 Years Classroom Programme

MARKS

715

720



AIR

45th

SHASHANK SHARMA
1 Year Classroom Programme

MARKS

715

720



AIR

60th

KRISHNAMURTI P S
Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

64th

SAM SHREYAS JOSEPH
2 Years Classroom Programme

MARKS

715

720



AIR

76th

DARSH PAGHDAR
Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

80th

RISHABH SHAH
Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

91st

SOHAM SHAMKANT GAYKAR
Online Crash Course
(Distance Learning Programme)

MARKS

715

720



AIR

103rd

D N NITHIN
4 Years Classroom Programme

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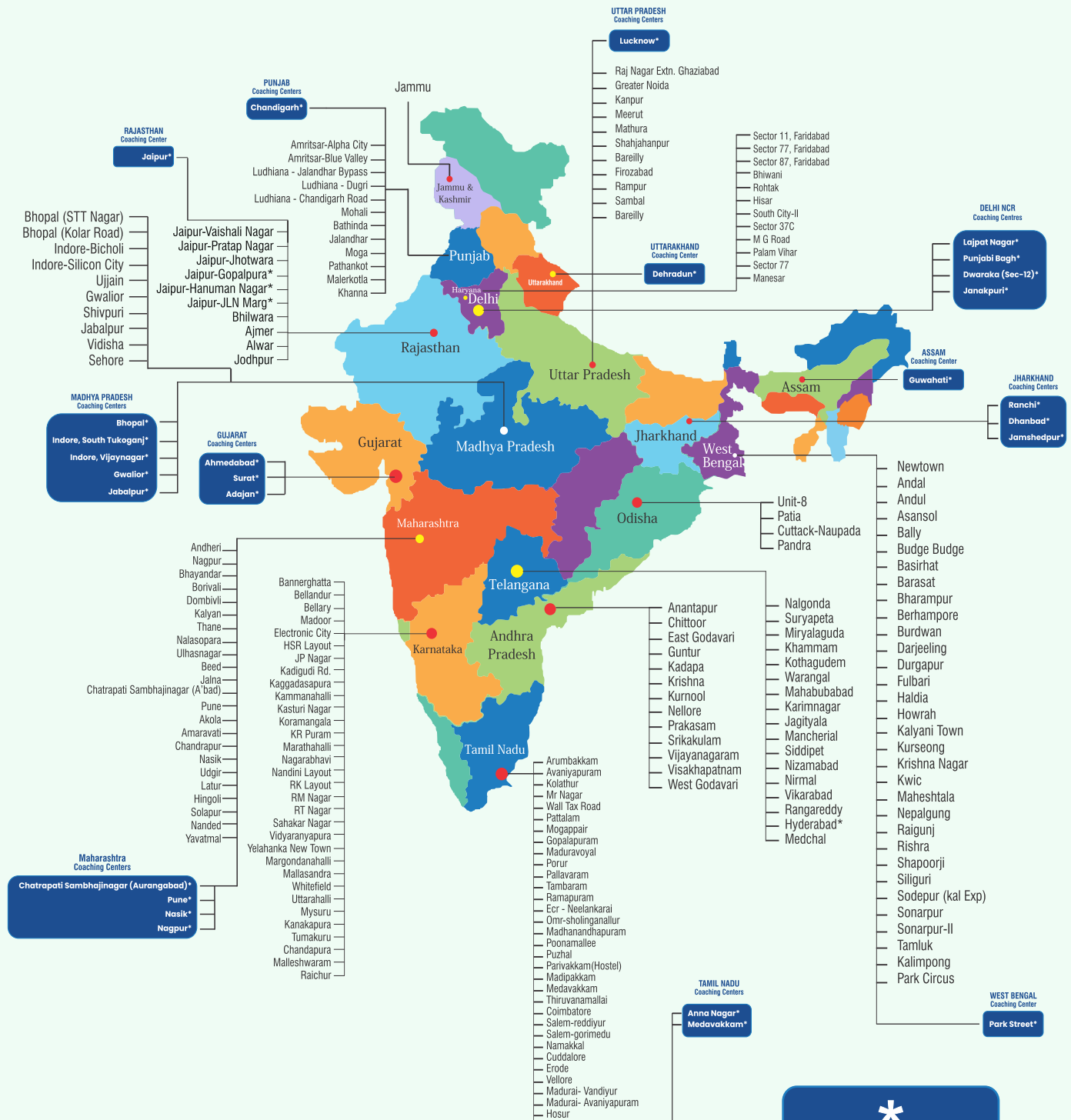
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