NARAYANA
IIT-JEE/NEET/FOUNDATION
JAIPUR CENTER

46 YEARS OF EXCELLENCE

SAMPLE PAPER - 6

NEET (UG) | 2025

Duration : 3 Hrs. | Maximum Marks : 720

ANSWER KEY

Q.	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
Ans.	1	3	4	1	4	4	3	4	1	1	1	3	3	3	1	1	1	3	3	4
Q.	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40
Ans.	2	4	4	3	2	3	4	3	2	2	1	4	3	4	3	3	1	1	3	1
Q.	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60
Ans.	2	3	2	4	4	1	1	2	4	3	3	4	3	1	4	2	4	3	4	4
Q.	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80
Ans.	3	1	4	1	1	1	2	1	3	2	3	2	4	4	2	3	4	2	2	4
Q.	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100
Ans.	1	2	4	2	4	2	3	4	4	2	4	2	2	1	4	2	3	2	3	3
Q.	101	102	103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120
Ans.	1	4	2	4	3	3	4	4	3	4	4	3	3	1	2	1	2	2	1	4
Q.	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136	137	138	139	140
Ans.	3	2	4	4	3	3	3	1	1	2	2	3	3	2	2	2	4	3	3	3
Q.	141	142	143	144	145	146	147	148	149	150	151	152	153	154	155	156	157	158	159	160
Ans.	4	4	4	4	1	4	2	4	3	1	4	4	3	2	4	2	4	1	4	1
Q.	161	162	163	164	165	166	167	168	169	170	171	172	173	174	175	176	177	178	179	180
Ans.	1	1	2	2	3	4	1	3	1	1	2	2	2	1	1	2	2	1	2	4

SOLUTION

PART – I : PHYSICS

1.

Ans. 1

Sol. of $\frac{2\pi}{\lambda}$

$$\frac{2\pi ct}{\lambda}=\frac{2\pi x}{\lambda}$$

Unit of (ct = x = λ = A)

2.

Ans. 3

Sol. $LC=\frac{(0.5)}{50}=0.01\text{mm}$

3.

Ans. 4

Sol. $\Delta W_a=0.1\text{gf}, W_a=10.0\text{gf}, \therefore W=10-5=5\text{gf}$

$$\frac{\Delta f}{f}=\left(\frac{0.1}{10}+\frac{0.2}{5}\right)(100)=5\%$$

4.

Ans. 1

Sol. Let their velocities will be qual at the instant t.

Given that $x_p(t) = at + bt^2$ and $x_Q(t) = ft - t^2$.

Hence, vp(t) = a + 2bt

As according to questions

$V_p(t) = v_Q(t)$

$$\therefore a+2bt=f-2t \quad \text{or } t=\frac{f-a}{2(1+b)}$$

5.

Ans. 4

Sol. $A\Rightarrow\frac{u^2}{4}-u^2=-2gh_1$

$$B\Rightarrow\frac{u^2}{9}-u^2=-2gh_2$$

$$C\Rightarrow\frac{u^2}{16}-u^2=-2gh_3$$

$$AB=\frac{u^2}{2g}\left[\frac{8}{9}-\frac{3}{4}\right]=\frac{u^2}{2g}.\frac{5}{36}$$

$$BC=\frac{u^2}{2g}\left[\frac{15}{16}-\frac{8}{9}\right]=\frac{u2}{2g}.\frac{7}{144}$$

$$\therefore \frac{AB}{BC}=\frac{5}{36}\times\frac{144}{7}=\frac{20}{7}$$

6.

Ans. 4

Sol. If a body is moving on a frictionless surface, then its total mechanical energy remains conserved.

According to the conservation of mechanical energy,

$$(TE)_{initial}=(TE)_{final}$$

$$\Rightarrow (KE)_i+(PE)_i=(KE)_f+(PE)_f$$

$$0+mgh=\frac{1}{2}mv_A^2+0$$

$$\Rightarrow gh=\frac{v_A^2}{2} \text{ or } h=\frac{v_A^2}{2g}\ldots\ldots \text{ (i)}$$

In order to complete the vertical circle, the velocity of the body at point A should be

$$v_A=v_{min}=\sqrt{5gR}$$

Where, R is the radius of the body.

Here, $R=\frac{AB}{2}=\frac{D}{2}$

$$\Rightarrow v_{min}=v_A=\sqrt{\frac{5}{2}gD}$$

Substituting the value of v_A in Eq. (i), we get

$$h=\frac{\left(\sqrt{\left(\frac{5}{2}gD\right)}\right)^2}{2g}=\frac{5gD}{2\times 2g}=\frac{5}{4}D$$

7.

Ans. 3

Sol. Initial momentum = Final momentum

$\therefore m_1v_1 + m_2v_2 = (m_1 + m_2)v$

Given, $v_1 = 36$ km/h $= 36 \times \frac{5}{18} = 10\,m/s$,

$v_2 = 0$

$m_1 = 2$ kg, $m_2 = 3$ kg

$\therefore v = \frac{m_1v_1 + m_2v_2}{m_1 + m_2} = \frac{2\times10 + 3\times0}{2+3}$

Or $v = \frac{20}{5} = 4\,m/s$

Loss in kinetic energy

$= \frac{1}{2}m_1v_1^2 + \frac{1}{2}m_2v_2^2 - \frac{1}{2}(m_1 + m_2)v^2$

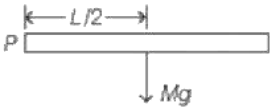
$= \frac{1}{2}\times2\times(10)^2 + 0 - \frac{1}{2}(2+3)\times(4)^2$

$=100 - 40 = 60$ J

9.

Ans. 1

Sol. Torque on the rod is equal to moment of weight of rod about P.



Torque on the rod = Moment of weight of the rod about P

$\tau = Mg\frac{L}{2}$...(i)

$\because$ Moment of inertia of rod

About $P = \frac{ML^2}{3}$...(ii)

As $\tau = I\alpha$

From Eqs. (i) and (ii), we get

$Mg\frac{L}{2} = \frac{ML^2}{3}\alpha \Rightarrow \alpha = \frac{3g}{2L}$

12.

Ans. 3

Sol. PT_2 = constant

$\Rightarrow \left(\frac{nRT}{V}\right)T^2 =$ constant

$\Rightarrow T^3 = KV$

$\Rightarrow \frac{dV}{dT} = \frac{3T^2}{K} = \frac{3V}{T}$

Hence $\gamma = \frac{1}{V}\left(\frac{dV}{dT}\right) = \frac{3}{T}$

14.

Ans. 3

Sol. $B_1 = \frac{\mu_0 i}{2R}$

$B_2 = \frac{\mu_0 iR^2}{2(R^2 + x^2)^{3/2}} = \frac{\mu_0 I}{2\times2\sqrt{2}R}$

$\frac{B_1}{B_2} = \frac{2\sqrt{2}}{1}$

16.

Ans. 1

Sol. Let resultant temperature is T°C

$\Rightarrow 10\times1(60^\circ - T) = 5\times1\times(T - 20^\circ)$

$\Rightarrow T = \left(\frac{140}{3}\right)^\circ C$

17.

Ans. 1

Sol. Energy required = (T . E) f – .(T . E)$_i$

$\Rightarrow E = \frac{-GMm}{8R} - \left(\frac{-GMm}{4R}\right)$

$\Rightarrow E = \frac{GMm}{8R} = 1.57\times10^9$ J

21.

Ans. 2

Sol. ${}^{237}_{93}Y \rightarrow {}^{229}_{89}Z + 2{}^4_2He +$ Energy

${}^A_ZX \rightarrow {}^{237}_{93}Y + {}^0_{+1}\beta + 2v$.

Z = 94 ; A=237

22.

Ans. 4

Sol. V = 6xy – y + 2yz

$$E = -\frac{\partial V}{\partial x}\hat{i} - \frac{\partial v}{\partial y}\hat{j} - \frac{\partial v}{\partial z}\hat{k}$$

$$= -6y\hat{i} - (6x - 1 + 2z)\hat{j} - (2y)\hat{k}$$

$$= -6\hat{i} - 5\hat{j} - 2\hat{k}$$

28.

Ans. 3

Sol. $i = \frac{V}{R} = \frac{10}{2} = 5A$

$$U = \frac{1}{2}Li^2 = \frac{1}{2} \times 2 \times 25 = 25J$$

31.

Ans. 1

Sol. Given, Young's double slit experiment, having two slits of width are in the ratio of 1: 25.

So, ratio of intensity,

$$\frac{I_1}{I_2} = \frac{W_1}{W_2} = \frac{1}{25} \Rightarrow \frac{I_2}{I_1} = \frac{25}{1}$$

$$\therefore \frac{I_{max}}{I_{min}} = \frac{\left(\sqrt{I_2} + \sqrt{I_1}\right)^2}{\left(\sqrt{I_2} - \sqrt{I_1}\right)^2} = \left(\frac{\sqrt{\frac{I_2}{I_1}} + 1}{\sqrt{\frac{I_2}{I_1}} - 1}\right)^2$$

$$\Rightarrow \left[\frac{5+1}{5-1}\right]^2 = \left(\frac{6}{4}\right)^2 = \frac{36}{16} = \frac{9}{4}$$

Thus, $\frac{I_{max}}{I_{min}} = \frac{9}{4}$

33.

Ans. 3

Sol. Excess energy of e^- appears as photon.

From Rydberg's formula,

$$\frac{1}{\lambda} = R\left(\frac{1}{n_f^2} - \frac{1}{n_i^2}\right) = R\left(\frac{1}{2^2} - \frac{1}{3^2}\right) = \frac{5R}{36}$$

$$\frac{1}{\lambda'} = R\left(\frac{1}{3^2} - \frac{1}{4^2}\right) = \frac{7R}{144}$$

$$\frac{1}{\lambda} / \frac{1}{\lambda'} = \frac{5R}{36} \div \frac{7R}{144}$$

$$\Rightarrow \frac{\lambda'}{\lambda} = \frac{5R}{36} \times \frac{144}{7R} = \frac{20}{7}$$

$$\Rightarrow \lambda' = \frac{20}{7}\lambda$$

37.

Ans. 1

Sol. $K_1 = 9K_2, l_1 = 18cm, l_2 = 6cm,$

$\theta_1 = 100^\circ C, \theta_2 = 0^\circ C$

Temperature of the junction

$$\theta = \frac{\frac{K_1}{l_1}\theta_1 + \frac{K_2}{l_2}\theta_2}{\frac{K_1}{l_1} + \frac{K_2}{l_2}}$$

$$\Rightarrow \theta = \frac{\frac{9K_2}{18}100 + \frac{K_2}{6} \times 0}{\frac{9K_2}{18} + \frac{K_2}{6}} = \frac{50+0}{8/12} = 75^\circ C$$

38.

Ans. 1

Sol. By adjoining graph $W_{AB} = 0$ and $W_{BC} = 8 \times 10^4 [5 - 2] \times 10^{-3} = 240$ J

$$\therefore W_{AC} = W_{AB} + W_{BC} = 0 + 240 = 240J$$

Now, $\Delta Q_{AC} = \Delta Q_{AB} + \Delta Q_{BC} = 600 + 200 = 800J$

From first law of thermodynamics

$$\Delta Q_{AC} = \Delta U_{AC} + \Delta W_{AC} \Rightarrow 800 = \Delta U_{AC} + 240$$

$$\Rightarrow \Delta U_{AC} = 560J$$

40.

Ans. 1

Sol. $\vec{a}_1 = 3\hat{i} + 3\hat{j} ms^{-2}, a_2 = 0$

$$\vec{v}_{cm} = \frac{m_1\vec{v}_1 + m_2\vec{v}_2}{m_1 + m_2} = \frac{v_1 + v_2}{2} = \frac{2\dot{i} + 2\dot{j}}{2}$$

$$= (\hat{i} + \hat{j})m/s$$

$$\vec{a}_{cm} = \frac{\vec{a}_1 + \vec{a}_2}{2} = \frac{3\hat{i} + 3\hat{j} + 0}{2} = \frac{3}{2}(\hat{i} + \hat{j})m/s^2$$

Now, $\vec{a}_{cm}$ and $\vec{v}_{cm}$ are along same direction.

$\therefore$ COM of system of particles move on straight line.

42.

Ans. 3

Sol. $B = \vec{B}_{xy} + \vec{B}_{yz} + \vec{B}_{zx}$

$\vec{B}_{xy} = \frac{\mu_0}{4\pi}\frac{I}{R}(\theta)\hat{k}$

$\vec{B}_{yz} = \frac{\mu_0}{4\pi}\frac{I}{R}\theta\hat{i}$

$\vec{B}_{zx} = \frac{\mu_0 I}{4\pi R}(\theta)\hat{j}$

$\vec{B} = \frac{\mu_0/\theta}{4\pi R}(\hat{k}+\hat{i}+\hat{j})$

$= \frac{\mu_0}{4\pi} \times \frac{4}{5\times10^{-2}}\frac{\pi}{2}(\hat{i}+\hat{j}+\hat{k})$

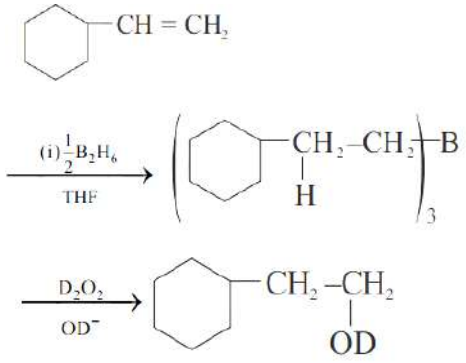
$= 10\mu_0(\hat{i}+\hat{j}+\hat{k})T$

PART – II : CHEMISTRY

48.

Ans. 2

Sol.



49.

Ans. 4

Sol. A is free radical substitution and B is E.S.R product.

50.

Ans. 3

Sol. According to stability of intermediate carbocation.

51.

Ans. 3

Sol. Emission spectrum is line spectrum.

52.

Ans. 4

Sol. Power of bulb $= 40\,W$

Energy emitted by a bulb = power $= 40\times20 = 800\,J$

Energy of photons emitted by bulb

$= 800\times\frac{80}{100} = 640\,J$; $E = \frac{nhc}{\lambda}$

$640 = \frac{n\times6.6\times10^{-34}\times3\times10^{8}}{620\times10^{-9}}$; $n = 2\times10^{21}$ photons

54.

Ans. 1

Sol. NO (Neutral)

Al_2O_3 (Amphoteric)

Na_2O (Basic)

Cl_2O_7 (Acidic)

57.

Ans. 4

Sol. Formaldehyde and benzaldehyde answer Cannizzaro reaction which involves the conversion of the aldehyde into a mixture of an alcohol and an acid. Thus option (4) can be chosen.

58.

Ans. 3

Sol. $C_6H_5COOH \xrightarrow[\Delta]{NH_3} C_6H_5CONH_2$

$\xrightarrow{NaOBr} C_6H_5NH_2 \xrightarrow[460\,K]{Conc.\ H_2SO_4}$

59.

Ans. 4

Sol. Carbohydrates $\rightarrow$ Glycosidic linkage

Proteins $\rightarrow$ Peptide linkage

Valine $\rightarrow$ Essential amino acid

Adenine $\rightarrow$ Purine base

60.

Ans. 4

Sol. $CH_3-CH_2-OH \rightarrow \underset{(A)}{CH_3CH_2I} \rightarrow CH_3CH_2MgI \rightarrow$

$CH_3\underset{\downarrow}{C}H_2CH_2OMgI$

$CH_3CH_2CH_2OH$

61.

Ans. 3

Sol. $C_6H_5N_2^+Cl^- \xrightarrow[Cu_2Cl_2]{HCl} C_6H_5Cl$ (Sand meyer rxn)

$C_6H_5N_2^+ + Cl^- \xrightarrow[Cu]{HBr} C_6H_5Br$ (Gattermann rxn),

$C_6H_5N_2^+Cl^- \xrightarrow{C_6H_5OH}$ (Coupling reaction)

$C_6H_5NH_2 \xrightarrow[0 \text{ to } 5^\circ C]{NaNO_2/HCl}$ (Diazotisation

62.

Ans. 1

Sol. Gabriel phthalimide synthesis is used to prepare aliphatic primary amines only

64.

Ans. 1

Sol. The correct order is $NH_3 > PH_3 > AsH_3 > SbH_3 > BiH_3$. It is due to extensive hydrogen bonding in NH_3 molecules.

65.

Ans. 1

Sol. Because reaction occurs, $\Delta G < 0$. Because of conversion of 2 moles of O_3 to 3 moles of O_2 the entropy is increasing, $\Delta S > 0$. The conversion of O_3 to O_2 is exothermic, $\Delta H < 0$.

66.

Ans. 1

Sol. Factual

69.

Ans. 3

Sol. $\Delta H = \Delta U + \left(P_2V_2 - P_1V_1\right)$

72.

Ans. 2

Sol. 188 g of AgBr contains 80 g bromine

0.376 g of AgBr contains $\frac{80}{188} \times 0.376\,g$

bromine

Percentage of bromine $= \frac{80 \times 0.376 \times 100}{188 \times 0.5}$

$= 32\%$

73.

Ans. 4

Sol. $CH_3CH_2CH_2OH \xrightarrow{SOCl_2} \underset{(A)}{CH_3CH_2CH_2Cl}$

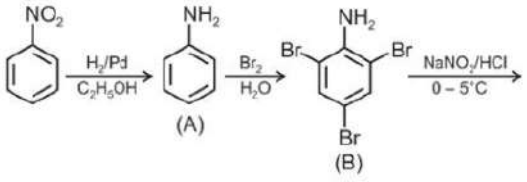
$\xrightarrow{KCN} \underset{(B)}{CH_3CH_2CH_2CN}$

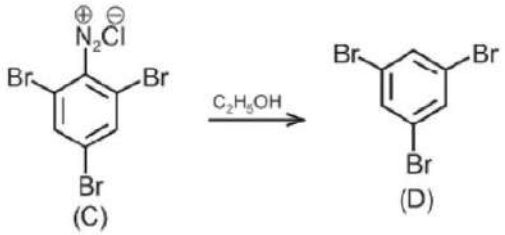
$\xrightarrow[(ii)H_2O]{(i)DIBAlH} \underset{(C)}{CH_3CH_2CH_2CHO}$

75.

Ans. 2

Sol.

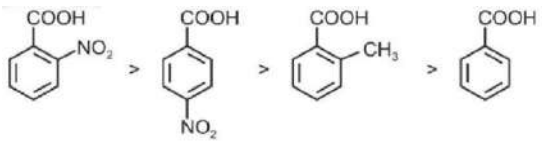




76.

Ans. 3

Sol. Correct order of acidic strength is



77.

Ans. 4

Sol. Higher the negative charge on the metal, stronger is back donation and stronger is metal carbon bond.

78.

Ans. 2

Sol. $\text{Solubility (s)} = \dfrac{1.95\times10^{-5}}{97.5} = 0.02\times10^{-5}$

$= 2\times10^{-7}\,\text{M}\ \text{Cu(OH)}_2(s) \rightleftharpoons \underset{s}{\text{Cu}^{2+}}(aq) + \underset{2s}{2\text{O}}(aq)$

$K_{sp} = \left[\text{Cu}^{2+}\right][\overline{\text{O}}\text{H}]^2 = s(2s)^2 = 4s^3$

$= 4\times\left(2\times10^{-7}\right)^3 = 3.2\times10^{-20}\,\text{mol}^3\,\text{L}^{-3}$

80.

Ans. 4

Sol. The given reaction is first order

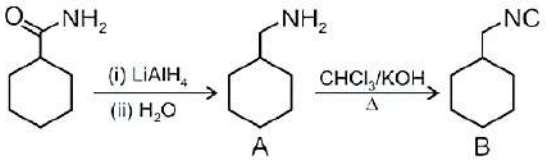
$t_{1/2} = \dfrac{0.693}{k} = \dfrac{0.693}{1.155\times10^{-3}} = 0.6\times10^3 = 600\,\text{s}$

$t_{1/2} = 10\,\text{min}.$

81.

Ans. 1

Sol.



82.

Ans. 2

Sol. $\Delta G = \Delta H - T\Delta S$ At equilibrium, $\Delta G = 0$

$\therefore 0 = \left(170\times10^3\,\text{J}\right) - \text{T}\left(170\text{JK}^{-1}\right) \therefore \text{T} = 1000\,\text{K}$

For spontaneity, ΔG is $-$ ve, which is possible

83.

Ans. 4

Sol. $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)}\text{CO}_{2(g)} + \text{H}_{2(g)}$

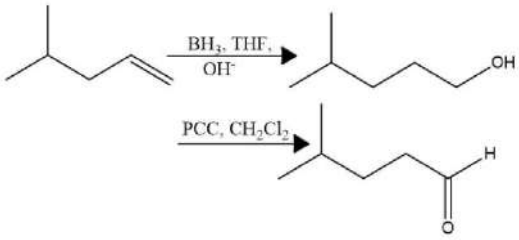
$K_1 = \dfrac{[\text{CO}_2][\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]}$ .....(l)

$\text{CH}_{4(g)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{(g)} + 3\text{H}_{2(g)}$

84.

Ans. 2

Sol.



85.

Ans. 4

Sol. NH_2OH is Hydroxylamine, so due to absence of carbon it is not able to give Prussian blue colour in Lassaigne's test.

86.

Ans. 2

Sol. $3\text{Cl}_2 + 6\text{NaOH} \rightarrow 5\text{NaCl} + \text{NaClO}_3 + 3\text{H}_2\text{O}$

88.

Ans. 4

Sol. HIO_4 oxidises $-\text{CH}_2\text{OH}$ to HCHO and breaks the $\text{C}-\text{C}$ bond of terminal CH_2OH groups

89.

Ans. 4

Sol. Vinyl / Aryl halides will not give AgBr ppt with AgNO_3 due to partial double bond character.

90.

Ans. 2

Sol. In presence of base inversion takes place.

46 YEARS OF EXCELLENCE

SRAMAYEVA JAYATHE

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NARAYANA JAIPUR STUDENTS DOMINATE IN 2025

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JEE (Main) 2025 Session-I

CITY TOPPER
JAIPUR
& RAJ. STATE TOPPER
100
Percentile

AYUSH SINGHAL
MPS
Regular Classroom

2nd CITY TOPPER
JAIPUR
99.99
Percentile

ARVIN GUPTA
Rishikul
Regular Classroom

GOLD MEDAL for India in **IJSO-2024**

International Junior Science Olympiad-2024,
Held at Bucharest ROMANIA (2 to 12-Dec., 2024)

BHAVYAA GUNWAL
Yearlong Classroom Student (CO-SPARK Batch)

International Olympiad (St-II) 2025

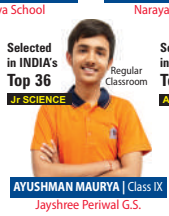
05 Selections for Stage-3
Orientation Cum Selection Camp (OCSC)

Indian National Olympiad (INO) 2025 Qualified
As per result declared by HBCSE & Nehru Science Center, NCSM on 05-03-2025

SAISHA JAIN | Class XI
Sarvodaya School

BHAVYAA GUNWAL | Class XI
Narayana E-Techno

NIDHI YADAV | Class XI
Srilal Convent



AYUSHMAN MAURYA | Class IX
Jayshree Periwal G.S.

DIVYANSH BANSAL | Class IX
Cambridge Court

Narayana Jaipur Center Toppers from the Very First Student Group

JEE (Adv) 2024

3 AIRs in Top 100 & 21 in Top 1000

AIR **33**

SPARSH GUPTA

AIR **49**

AVADH HINDOCHA

AIR **66**

MRIGANK GOEL

Further AIRs in Top 1000 : 116 (Prabuddha Sinha), 149 (Visharad Srivastava), 161(Mayukh Chowdhury), 168 (Abubakar Siddique), 303 (Shrey Nayakpara), 318 (Srutarshi Tripathi), 339 (Yash Vashisht), 354 (Aradhana R), 360 (Kathan Shah), 375 (Pratham Srivastava), 384 (Tushar), 448 (Vivaan Goswami), 543 (Shaurya Pratap), 569 (Akshat Khandelwal), 591 (Shourya Agarwal), 809 (Sagarika Sinha), 881 (Ishwin Kumar), 973 (Muthu S)

(* All these are Narayana Jaipur Center's **InfinITy SRG** batch students who attended online night classes by top faculties to boost rank in **JEE (Adv) 2024**)

NEET 2024

3 AIRs in Top 100 & 15 in Top 1000

AIR **37**
715/720

TANISHAK YADAV

AIR **45**
715/720

SHASHANK SHARMA

AIR **64**
715/720

SAM SHREYAS JOSEPH

Further AIRs in Top 1000 : 181(Atishay Jain), 316 (Krish Jain), 340 (Somya Garg), 350 (Dhruv Awasthy), 380 (Akshat Mishra), 562 (Ashvin Sharma), 606 (Purvika Singh), 616 (Shreya Jain), 666 (Surya Charan D), 681(Tanmay Singhal), 786 (Amogh), 966 (Chhavi Jain)
(All Scored 700+)

(* All these are Narayana Jaipur Center's Regular Classroom students of **RT-720 Course** (Nov'23 to May'24) for **NEET 2024**)

This very first student group result has proven the power of Top Level Mentorship, Genius Faculties and Excellent System at Narayana Jaipur Center

For any enquiry or assistance, call **0141-4848000**

www.narayanajaipur.com | ntsc.narayanatalent.com | jaipur@narayanagroup.com | **SUNDAY OPEN**

BATCH STARTING DATES

Moving to **Class 6, 7, 8, 9** & **10**
Starting: **02** & **09-Apr**, 2025

Moving to **Class 11* (JEE/NEET)**
Starting: **26-Mar** & **02-Apr**, 2025

Moving to **Class 12* (JEE/NEET)**
Starting: **19-Mar**, 2025

Appeared-in/Passed Class 12 (NEET)
Starting: **09** & **16-Apr**, 2025
12वीं पास हिंदी माध्यम **NEET** का पृथक बैच **11**-Jun से

Appeared-in/Passed Class 12 (JEE)
Starting: **23-Apr** & **07-May**, 2025

* Separate Batches for JEE & NEET as per New Education Policy (NEP)